

Assessment

Chapter Test B**Chapter: Chemical Formulas and Chemical Compounds**

PART I In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. Changing a subscript in a correctly written chemical formula will
- change the electron configuration of that element.
 - change the charges on the other ions in the compound.
 - change the formula so that it no longer represents the same compound.
 - have no effect on the formula.
- _____ 2. Using the Stock system of nomenclature, $\text{Cr}_2(\text{SO}_4)_3$ is named
- chromium(II) sulfate.
 - chromic sulfate.
 - dichromium trisulfate.
 - chromium(III) sulfate.
- _____ 3. In a polyatomic ion, the algebraic sum of the oxidation numbers of all atoms is equal to
- 0.
 - 10.
 - the number of atoms in the ion.
 - the charge on the ion.
- _____ 4. The first part of the name of a binary ionic compound is the
- cation.
 - polyatomic ion.
 - oxyanion.
 - anion.
- _____ 5. The empirical formula may not represent the actual composition of a(n)
- ionic compound.
 - crystal.
 - salt.
 - molecular compound.
- _____ 6. The formula for phosphoric acid is
- HPO_4
 - $\text{H}(\text{PO}_4)_2$
 - H_2PO_3
 - H_3PO_4

Chapter Test B Chemical Formulas And Compounds

Mcdougal Littell



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