

## Chapter 6 – Study Guide

Test Date: Friday, December 19, 2014

Use the following to study:

Textbook – Ch. 6

Ch. 6 homework packet

### Chemical Bonding:

- **Four types of bonds:**
  - *Nonpolar Covalent Bond* ( $\Delta EN$  (difference in electronegativity) = < 0.3)
    - EQUAL SHARING of electrons between the 2 bonding nonmetal atoms
  - *Polar Covalent Bond* ( $\Delta EN = 0.4-1.7$ )
    - UNEQUAL SHARING of electrons between the 2 bonding nonmetal atoms (because EN is much higher in one atom, the electrons are pulled closer to the atom in the bond with the higher EN value – like an uneven tug of war)
  - *Ionic Bond* ( $\Delta EN = > 1.7$ )
    - Difference in the EN values for the 2 bonding atoms (metal & nonmetal) is so large that the atom with the lower EN value ends up completely GIVING UP/TRANSFERRING ELECTRONS to the atom with the greater EN value
    - Ionic bond results due to the attraction between the POSITIVE cation (formed by losing electrons) and the NEGATIVE anion (formed by gaining electrons)
  - *Metallic Bond*
    - Bonding between two metallic atoms which occurs due to the attraction between a “sea of electrons” and the nuclei of bonded metal atoms
- \* Must memorize/know the difference in electronegativity ranges for each type of bond
  - \*\* Note: you should still know the definition of electronegativity (EN)
- **Properties of the Four Types of Chemical Bonds**
  - See the table at the bottom of Packet pg.5
  - You MUST be able to compare all four examples – NaCl (ionic), Sucrose (Polar), Stearic Acid/Wax (nonpolar), and Aluminum & Mercury (metallic) – in terms of melting point, conductivity of solid, solubility in water, conductivity in water (when applicable), solubility in hexane, hardness (ease with which it is scratched), and brittleness
- **Why & How do chemical bonds form?**
  - Why → because all atoms want to be at their **LOWEST POTENTIAL ENERGY**; for nearly all atoms (except the noble gases) this happens only when an atom chemically bonds to other atoms!!!
  - How → **ATTRACTIVE** and **REPULSIVE** forces
    - *Attractive forces* – positively charged nucleus of one atom is attracted to electrons in another atom nearby (because they have opposite charges)
    - *Repulsive forces* – 2 nuclei or 2 sets of electron clouds from nearby atoms repel one another (because they have the same charge)
  - **Bond Length**
    - ideal distance between 2 atoms that gives the lowest potential energy of the atoms by balancing out attractive & repulsive forces)
  - **Bond Energy**
    - Energy required to break apart a chemical bond (the higher the bond energy, the stronger the bond)

# Chemistry Chapter 6 Study Guide Answers

**Pamela Mills, Amina El-Ashmawy**



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web adapun jenis jenis motif anyaman yang unik diantaranya yaitu kelabang geometris ceplok antik segitiga dan liris motif corak anyaman rotan sangat beragam dirancang oleh para perajin rotan secara turun temurun berdasarkan tradisi nenek moyang bila ada anyaman rotan kreasi baru itu hanya pengembangan dari motif anyaman yang sudah ada

motif anyaman sederhana macam macam motif anyaman blogger - Feb 09 2023

web jan 26 2016 keindahan kerajinan anyaman ditentukan oleh motif motif hias yang dihasilkan dari variasi anyaman jenis bahan untuk membuat kerajinan tangan anyaman dikelompokkan menjadi 2 yaitu a bahan alam daun kelapa mendong daun lontar bambu pandan rotan b bahan buatan benang kertas lembaran plastik corak corak

*gambar motif anyaman png dan vektor dengan background* - Jan 28 2022

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sejarah fungsi dan berbagai jenis kerajinan anyaman - Apr 30 2022

web jan 19 2022 anyaman biasanya digunakan untuk barang sehari hari seperti aseupan pengukus nasi boboko tempat nasi besek kemasan hantaran hihid kipas samak tikar serta keranjang anyaman yang digunakan sebagai barang sehari hari umumnya anyaman halus dan motifnya lebih terlihat