

ANSWERS TO CONVERSION FACTOR PROBLEMS

Conversion factors can be used to convert units or to convert between equivalent ways of expressing a quantity. The quantity in the problem is multiplied by one or more "conversion factors," in which the numerator is equal to the denominator. Since the numerator and denominator of the conversion factor are equal, multiplying by the conversion factor is like multiplying by 1 and thus does not change the value of the original quantity. Use the table of English to Metric equivalents as needed. All answers should be in *significant figures*!

A) Problems with a single conversion factor.

1. How many cm are in 18.9 inches?

$$18.9 \text{ in} \times \frac{2.54 \text{ cm}}{1 \text{ in}} = 47.5 \text{ cm}$$

2. How many grams are in 0.143 ounces?

$$0.143 \text{ oz} \times \frac{1 \text{ gram}}{0.03527 \text{ oz}} = 4.05 \text{ grams}$$

B) Problems with two or more conversion factors.

1. How many mL are in 0.037 quarts?

$$0.037 \text{ qt} \times \frac{1 \text{ L}}{1.057 \text{ qt}} \times \frac{1000 \text{ mL}}{1 \text{ L}} = 35 \text{ mL}$$

2. How many grams are in 0.397 pounds (lbs)?

$$0.397 \text{ lb} \times \frac{1 \text{ kg}}{2.20 \text{ lb}} \times \frac{1000 \text{ g}}{1 \text{ kg}} = 180 \text{ g} = 1.80 \times 10^3 \text{ g}$$

3. How many micrograms are in 6.8×10^{-7} ounces?

$$6.8 \times 10^{-7} \text{ oz} \times \frac{1 \text{ g}}{0.0353 \text{ g}} \times \frac{1000 \text{ mg}}{1 \text{ g}} \times \frac{1000 \text{ } \mu\text{g}}{1 \text{ mg}} = 19 \text{ } \mu\text{g}$$

C) Using density as a conversion factor. The density of Al is 2.70 g/cm^3 .

Remember, $1 \text{ mL} = 1 \text{ cm}^3$.

1. What is the volume of 0.810 grams of Al?

$$0.810 \text{ g} \times \frac{1 \text{ cm}^3}{2.70 \text{ g}} = 0.300 \text{ cm}^3 \text{ (or } 0.300 \text{ mL)}$$

2. Find the mass in grams of 0.327 liters of Al.

$$0.327 \text{ L} \times \frac{1000 \text{ mL}}{1 \text{ L}} \times \frac{1 \text{ cm}^3}{1 \text{ mL}} \times \frac{2.70 \text{ g}}{1 \text{ cm}^3} = 883 \text{ g}$$

Chemistry Conversion Factors Practice Problems With Solutions

Ralph A. Burns, John William Hill



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General Chemistry, Reactions First Kevin Revell, 2024-12-04 Revell's General Chemistry empowers students to grasp essential topics and concepts with more ease Using a friendly approach the text uses metaphors and relatable examples to demystify even the most challenging subjects in general chemistry

A Learning Program for Chemistry Peter Demmin, 1993-11 *Improving Student Comprehension in Chemistry Laboratories* Tracy Lynn Haroff, 2006

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Essentials of Chemistry Ralph A. Burns, John William Hill, 1995 The second edition of this chemistry textbook that uses practice examples and applications relating chemistry to our lives and the environment

New Technical Books New York Public Library, 1969

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