

CHEMISTRY

OPEN

PERIODIC TABLE OF THE ELEMENTS

H — atomic number
1.008 — atomic weight



ATOMIC STRUCTURE

Atomic Number, Z , of an element is the number of protons in the nucleus. For neutral atoms, $Z = +$ the number of electrons. The mass number, A , is the sum of the number of protons and the number of neutrons. Isotopes of an element have the same Z , different A . The atomic mass is the weighted average of isotopes. See also: Atomic Mass, Atomic Weight, Crystal Structure of Solids, Isotopes for a given element.

ATOMIC QUANTUM NUMBERS AND ORBITALS

Principal Quantum Number (n): $n = 1, 2, 3, \dots$
Angular Momentum Quantum Number (l): $l = 0, 1, 2, \dots, n-1$
Magnetic Quantum Number (m_l): $m_l = -l, -(l-1), \dots, 0, \dots, (l-1), l$
Spin Quantum Number (m_s): $m_s = +\frac{1}{2}, -\frac{1}{2}$

NUCLEAR CHEMISTRY

Nuclear reactions alter the number, particles, energy, or mass of a nucleus. Energy is released or absorbed. The mass number, A , and atomic number, Z , are conserved. The mass defect, Δm , is the difference between the mass of the nucleus and the sum of the masses of its constituent nucleons. The binding energy, B , is the energy required to disassemble a nucleus into its constituent nucleons. The binding energy per nucleon, B/A , is a measure of the stability of a nucleus. The mass defect, Δm , is the difference between the mass of the nucleus and the sum of the masses of its constituent nucleons. The binding energy, B , is the energy required to disassemble a nucleus into its constituent nucleons. The binding energy per nucleon, B/A , is a measure of the stability of a nucleus.

MANY-ELECTRON ATOMS

Orbitals and quantum numbers are used to describe the behavior of electrons in many-electron atoms and ions. A many-electron atom has many different energy levels and orbital properties. The main result is that orbitals have different energies. The filling of levels is guided by the Aufbau Principle.

n	l	m_l	Orbitals
1	0	0	1s
2	0	0	2s
2	1	0	2p _z
2	1	±1	2p _x , 2p _y
3	0	0	3s
3	1	0	3p _z
3	1	±1	3p _x , 3p _y
3	2	0	3d _{z²}
3	2	±1	3d _{xz} , 3d _{yz}
3	2	±2	3d _{x²-y²} , 3d _{xy}

Atomic configuration: How orbitals are filled. Pauli Exclusion Principle: Each orbital has a unique set of quantum numbers. Hund's Rule: Electrons fill orbitals singly, with parallel spins, before pairing up. Aufbau Principle: Orbitals are filled in order of increasing energy.



Block	Group	Element	Symbol	Atomic Number		
s-block	1	H	H	1		
		Li	Li	3		
	2	Be	Be	4		
		Mg	Mg	12		
	p-block	13	B	B	5	
			Al	Al	13	
		14	C	C	6	
			Si	Si	14	
		15	N	N	7	
			P	P	15	
		16	O	O	8	
			S	S	16	
		17	F	F	9	
			Cl	Cl	17	
		18	Ne	Ne	10	
			Ar	Ar	18	
		d-block	3	Sc	Sc	21
				Ti	Ti	22
4	V		V	23		
	Cr		Cr	24		
5	Mn		Mn	25		
	Fe		Fe	26		
6	Co		Co	27		
	Ni		Ni	28		
7	Cu		Cu	29		
	Zn		Zn	30		
f-block	8	La	La	57		
		Ce	Ce	58		
	9	Pr	Pr	59		
		Nd	Nd	60		
	10	Pm	Pm	61		
		Sm	Sm	62		
	11	Eu	Eu	63		
		Gd	Gd	64		
	12	Tb	Tb	65		
		Dy	Dy	66		
13	Ho	Ho	67			
	Er	Er	68			
14	Tm	Tm	69			
	Yb	Yb	70			
15	Lu	Lu	71			
	Hf	Hf	72			

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JR Anderson



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