

Chemistry 338 2nd Semester Final Exam Review Sheet

Unit 1 - Matter & Measurement – Chapters 1 & 2

1. A student experimentally determines the density of an object to be 4.2 g/mL. If the accepted value is 5.0 g/mL, calculate the % error.

$$\% \text{ error} = \frac{\text{experimental} - \text{accepted}}{\text{accepted}} \times 100$$

$$\% \text{ error} = \frac{4.2 - 5.0}{5.0} \times 100 = -16\% \text{ error}$$

Unit 2 – Periodic Table – Chapter 5

Element	Symbol	Atomic #	Protons	Neutrons	Electrons	Mass #	Charge
Tungsten ion	$^{184}_{74}\text{W}^{+2}$	74	74	110	72	184	+2

3. Using the following atomic radius values, predict the atomic radius of potassium (K):
Li – 145 pm, Na – 180 pm, K - ?, Rb – 235 pm
K – 220 pm

Unit 3 – Reactions/Compounds – Chapters 7 & 8

4. Classify the compound as ionic or covalent. Name the compound
- Fe_2O_3 - iron (III) oxide (ionic - roman numerals for transition metals, change ending to -ide)
 - Cl_2O_8 - dichlorine octaoxide (covalent - use prefixes, change ending to -ide)
5. Classify the compound as ionic or covalent. Write the formula
- Trisulfur Dinitride – S_3N_2 (covalent)
 - Zinc Hydroxide – $\text{Zn}(\text{OH})_2$ (ionic - look up ion charges and criss-cross) $\text{Zn}^{+2} \text{OH}^{-1}$
 - Chromium (III) nitrate – $\text{Cr}(\text{NO}_3)_3$ (ionic - criss-cross) $\text{Cr}^{+3} \text{NO}_3^{-2}$
6. Write the equation, balance it, and classify the type of reaction:
Iron (II) sulfate reacts with sodium hydroxide to yield iron(II) hydroxide and sodium sulfate
- $$1 \text{ FeSO}_4 + 2 \text{ NaOH} \rightarrow 1 \text{ Fe}(\text{OH})_2 + 1 \text{ Na}_2\text{SO}_4$$
- double displacement (two ionic compounds reacting)
7. Balance the equation and classify the type of reaction:
- $$1 \text{ Cu} + 2 \text{ AgNO}_3 \rightarrow 1 \text{ Cu}(\text{NO}_3)_2 + 2 \text{ Ag}$$
- Single displacement (element and a compound reacting)

Unit 4 – Moles – Chapter 3

8. Calculate the molar mass of $\text{Fe}_2(\text{SO}_4)_3$.
- $$\begin{aligned} \text{Fe } (2)(55.9 \text{ g/mol}) &= 111.8 \text{ g/mol} \\ \text{S } (3)(32.1 \text{ g/mol}) &= 96.3 \text{ g/mol} \\ + \text{O } (12)(16.0 \text{ g/mol}) &= +192 \text{ g/mol} \\ &= 400.1 \text{ g/mol} \end{aligned}$$

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A Loxley



Chemistry Spring Semester Final Exam Review Sheet:

Survival Handbook for the New Chemistry Instructor Diane M. Bunce, Cinzia M. Muzzi, 2004 This book provides an overview of the issues facing new chemistry faculty in preparation for teaching Serving as a reference to answer specific questions new chemistry faculty encounter this book is comparable to sitting down with a colleague in the department and talking through some ideas or gaining some pointers on how to avoid common pitfalls It is the one single place new chemistry faculty can go to find practical information on how to teach and how to prepare for teaching their first course Chapters are written both by established experts in the field and by new professors within their first couple of years of teaching

Cornell University Courses of Study Cornell University, 1992 Congressional Record United States. Congress, 1971 The Congressional Record is the official record of the proceedings and debates of the United States Congress It is published daily when Congress is in session The Congressional Record began publication in 1873 Debates for sessions prior to 1873 are recorded in The Debates and Proceedings in the Congress of the United States 1789 1824 the Register of Debates in Congress 1824 1837 and the Congressional Globe 1833 1873 **Portable Technologies** Robert Tinker, Joseph Krajcik, 2012-12-06 Education has traditionally studied the world by bringing it into the classroom This can result in situated learning that appears to students to have no relevance outside the classroom Students acquire inert decontextualized knowledge that they cannot apply to real problems The obvious solution to this shortcoming is to reverse the situation and bring the classroom to the phenomena to learn in a rich real world context The problem with the real world is that it is complex and filled with interactions that are hard to sort out The editors and authors believe that the right tools might help students with this sorting process and result in learning in rich contexts This book is an account of a series of experiments designed to explore the validity of this insight **The Journal of Materials Education** ,1998 *The Virginia Journal of Science* ,2005 **Virginia Journal of Science** Ruskin Skidmore Freer, 2006 Getting Into Medical School Donald J. Solomon, 1979 *School* ,1929-09 **Printers' Ink** ,1957 The New Statesman and Nation ,1944 Final Exam Review Sheet ,2002 *Chemistry and Student Study Guide and Study Guide, Fifth Edition and Smarthinking and Letter* Zumdahl, 2002-01-01 *Fast Track: Chemistry* The Princeton Review, 2020-12-08 GET UP TO SPEED WITH FAST TRACK CHEMISTRY Covering the most important material taught in high school chem class this essential review book breaks need to know content into accessible easily understood lessons Inside this book you ll find Clear concise summaries of the most important concepts terms and functions in chemistry Diagrams charts and graphs for quick visual reference Easy to follow content organization and illustrations With its friendly straightforward approach and a clean modern design crafted to appeal to visual learners this guidebook is perfect for catching up in class or getting ahead on exam review Topics covered in Fast Track Chemistry include Atomic structure Covalent bonding Intermolecular forces Stoichiometry Precipitation reactions Gas laws Thermochemistry Equilibrium and the solubility product constant Redox reactions Electrochemistry Acids and bases

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