

## Chapter 6 – Study Guide

Test Date: Friday, December 19, 2014

Use the following to study:

Textbook – Ch. 6

Ch. 6 homework packet

### Chemical Bonding:

- **Four types of bonds:**
  - *Nonpolar Covalent Bond* ( $\Delta EN$  (difference in electronegativity) =  $< 0.3$ )
    - **EQUAL SHARING** of electrons between the 2 bonding nonmetal atoms
  - *Polar Covalent Bond* ( $\Delta EN = 0.4-1.7$ )
    - **UNEQUAL SHARING** of electrons between the 2 bonding nonmetal atoms (because EN is much higher in one atom, the electrons are pulled closer to the atom in the bond with the higher EN value – like an uneven tug of war)
  - *Ionic Bond* ( $\Delta EN = > 1.7$ )
    - Difference in the EN values for the 2 bonding atoms (metal & nonmetal) is so large that the atom with the lower EN value ends up completely **GIVING UP/TRANSFERRING ELECTRONS** to the atom with the greater EN value
    - Ionic bond results due to the attraction between the **POSITIVE** cation (formed by losing electrons) and the **NEGATIVE** anion (formed by gaining electrons)
  - *Metallic Bond*
    - Bonding between two metallic atoms which occurs due to the attraction between a “sea of electrons” and the nuclei of bonded metal atoms
- \* Must memorize/know the difference in electronegativity ranges for each type of bond
  - \*\* Note: you should still know the definition of electronegativity (EN)
- **Properties of the Four Types of Chemical Bonds**
  - See the table at the bottom of Packet pg.5
  - You **MUST** be able to compare all four examples – NaCl (ionic), Sucrose (Polar), Stearic Acid/Wax (nonpolar), and Aluminum & Mercury (metallic) – in terms of melting point, conductivity of solid, solubility in water, conductivity in water (when applicable), solubility in hexane, hardness (ease with which it is scratched), and brittleness
- **Why & How do chemical bonds form?**
  - **Why** → because all atoms want to be at their **LOWEST POTENTIAL ENERGY**; for nearly all atoms (except the noble gases) this happens only when an atom chemically bonds to other atoms!!!
  - **How** → **ATTRACTIVE** and **REPULSIVE** forces
    - *Attractive forces* – positively charged nucleus of one atom is attracted to electrons in another atom nearby (because they have opposite charges)
    - *Repulsive forces* – 2 nuclei or 2 sets of electron clouds from nearby atoms repel one another (because they have the same charge)
  - **Bond Length**
    - ideal distance between 2 atoms that gives the lowest potential energy of the atoms by balancing out attractive & repulsive forces)
  - **Bond Energy**
    - Energy required to break apart a chemical bond (the higher the bond energy, the stronger the bond)

# Chapter 6 Study Guide Chemistry

**R Sandford**



## **Chapter 6 Study Guide Chemistry:**

## Whispering the Secrets of Language: An Emotional Quest through **Chapter 6 Study Guide Chemistry**

In a digitally-driven world wherever screens reign great and immediate conversation drowns out the subtleties of language, the profound strategies and psychological nuances hidden within phrases usually move unheard. Yet, located within the pages of **Chapter 6 Study Guide Chemistry** a fascinating fictional value pulsating with fresh thoughts, lies a fantastic journey waiting to be undertaken. Written by an experienced wordsmith, that wonderful opus attracts readers on an introspective journey, lightly unraveling the veiled truths and profound influence resonating within the very fabric of every word. Within the mental depths with this moving review, we can embark upon a genuine exploration of the book's core themes, dissect its interesting writing type, and yield to the effective resonance it evokes heavy within the recesses of readers' hearts.

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