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**Antonio Chella,Angelo  
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web determine the final temperature and state of the ammonia once the system reaches equilibrium assuming that no heat is lost to the surroundings use the information below normal boiling point of  $\text{NH}_3$   $33.3^\circ\text{C}$   $\Delta H_{\text{vap}} = 23.4 \text{ kJ mol}^{-1}$  specific heat  $\text{NH}_3$   $1.474 \text{ J g}^{-1} \text{ K}^{-1}$  specific heat  $\text{NH}_3$   $2.06 \text{ J g}^{-1} \text{ K}^{-1}$

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web apr 3 2023 the experiment described above can be summarized in a graph called a heating curve figure below figure 11 16 1 in the heating curve of water the temperature is shown as heat is continually added changes of state occur during plateaus because the temperature is constant

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web the cooling curve a plot of temperature versus cooling time in figure pageindex 4 plots temperature versus time as a  $75 \text{ g}$  sample of steam initially at  $1 \text{ atm}$  and  $200^\circ\text{C}$  is cooled although we might expect the cooling curve to be the mirror image of the heating curve in figure pageindex 3 the cooling curve is not an

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web once all of the material has solidified at the time indicated by point c the cooling will continue at a rate determined by the heat capacities of the two solids a and b the composition and of course the geometry of the experimental set up by measuring cooling curves for samples of varying composition one can map the entire phase

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