

Covalent Bonding Review Sheet – Dec 12, 2012

Name _____ class period _____

- Given the formula, a compound can be named using conventional systems that include Greek prefixes and Roman numeral where appropriate.
- Differences in electronegativity values can be used to predict where a bond fits on the continuum between ionic and covalent bonds.
- Since most compounds contain multiple bonds, a substance may contain more than one type of bond.

Section 1: Naming and Formula Writing

Write the formulas for the following covalent compounds

1. Nitrogen Monoxide NO
2. Dinitrogen Monoxide N₂O
3. Nitrogen Dioxide NO₂
4. Dinitrogen Trioxide N₂O₃
5. Dinitrogen Tetraoxide N₂O₄
6. Dinitrogen Pentaoxide N₂O₅
7. carbon dioxide CO₂
8. phosphorus triiodide PI₃
9. sulfur dichloride SCl₂
10. nitrogen trifluoride NF₃

Write names from the following formulas

11. N₂F₄ dinitrogen tetrafluoride
12. SCl₄ sulfur tetrachloride
13. ClF₃ chlorine trifluoride
14. SiO₂ silicon dioxide
15. P₄O₁₀ tetraphosphorus decaoxide
16. SiF₄ silicon tetrafluoride
17. H₂O dihydrogen monoxide
18. Br₂ dibromide or bromine
19. HBr hydrogen monobromide
20. N₂S₃ dinitrogen trisulfide

Concept Review Section Covalent Bonds Answer Key

S Marginson



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