

Chemistry Notes for class 12 Chapter 2 Solutions

Solution is a homogeneous mixture of two or more substances in same or different physical phases. The substances forming the solution are called components of the solution. On the basis of number of components a solution of two components is called binary solution.

Solute and Solvent

In a binary solution, solvent is the component which is present in large quantity while the other component is known as solute.

Classification of Solutions

(A) Following types of solutions are seen on the basis of physical state of solute and solvent.

S.No.	Solute	Solvent	Examples
Solid solutions			
1	Solid	Solid	Alloys
2	Liquid	Solid	Hydrated salts, Amalgam of Hg with Na
3	Gas	Solid	Dissolved gases in mineral
Liquid solutions			
4	Solid	Liquid	Salt/sugar solution in water
5	Liquid	Liquid	Alcohol in water
6	Gas	Liquid	Aerated drinks, O_2 in water
Gaseous solutions			
7	Solid	Gas	Iodine vapour in air
8	Liquid	Gas	Water vapour in air
9	Gas	Gas	Air ($O_2 + N_2$)

[if water is used as a solvent, the solution is called aqueous solution and if not, the solution is called non-aqueous solution.]

(B) Depending upon the amount of solute dissolved in a solvent we have the following types of solutions:

(i) **Unsaturated solution** A solution in which more solute can be dissolved without raising temperature is called an unsaturated solution.

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Table of Contents Chapter 25 Assessment Chemistry Answers

1. Understanding the eBook Chapter 25 Assessment Chemistry Answers
 - The Rise of Digital Reading Chapter 25 Assessment Chemistry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 25 Assessment Chemistry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 25 Assessment Chemistry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 25 Assessment Chemistry Answers
 - Personalized Recommendations
 - Chapter 25 Assessment Chemistry Answers User Reviews and Ratings

- Chapter 25 Assessment Chemistry Answers and Bestseller Lists
- 5. Accessing Chapter 25 Assessment Chemistry Answers Free and Paid eBooks
 - Chapter 25 Assessment Chemistry Answers Public Domain eBooks
 - Chapter 25 Assessment Chemistry Answers eBook Subscription Services
 - Chapter 25 Assessment Chemistry Answers Budget-Friendly Options
- 6. Navigating Chapter 25 Assessment Chemistry Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 25 Assessment Chemistry Answers Compatibility with Devices
 - Chapter 25 Assessment Chemistry Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 25 Assessment Chemistry Answers
 - Highlighting and Note-Taking Chapter 25 Assessment Chemistry Answers
 - Interactive Elements Chapter 25 Assessment Chemistry Answers
- 8. Staying Engaged with Chapter 25 Assessment Chemistry Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 25 Assessment Chemistry Answers
- 9. Balancing eBooks and Physical Books Chapter 25 Assessment Chemistry Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 25 Assessment Chemistry Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 25 Assessment Chemistry Answers
 - Setting Reading Goals Chapter 25 Assessment Chemistry Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 25 Assessment Chemistry Answers
 - Fact-Checking eBook Content of Chapter 25 Assessment Chemistry Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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