

Chapter 25: NUCLEAR CHEMISTRY

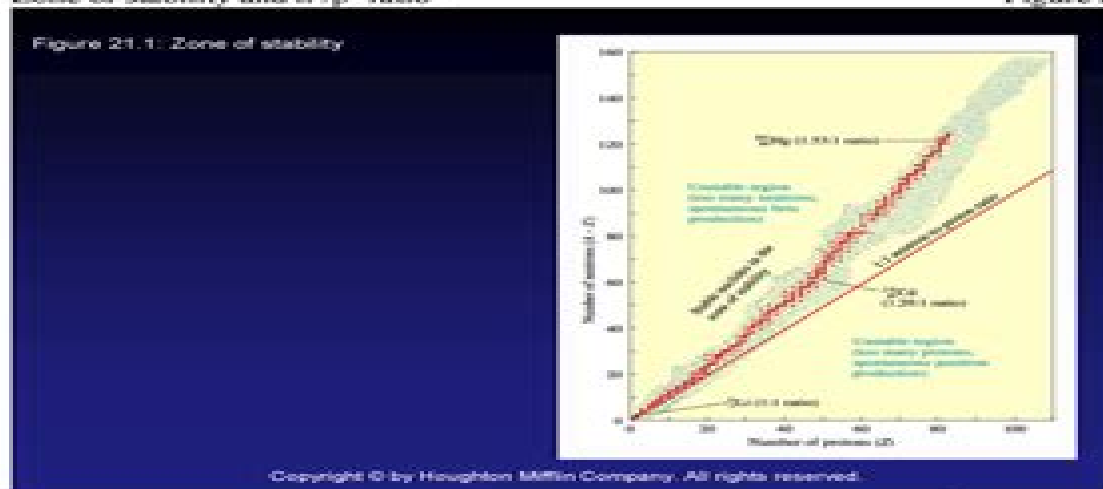
The number and arrangement of its electrons dominate the chemical reactivity and physical properties of an atom.

Hence, we focus on electron configurations in chemistry.

However, we live in a world with nuclear power plants, nuclear weapons, and medical applications of radioactive isotopes, including waste from their use and production.

Nuclear radiation can break chemical bonds, including bonds in DNA.

§25-7 Nuclear stability

Zone of stability and n^0/p^+ ratioFigure 25-7: Zumdahl 4th Fig 21.1

neutrons and protons are attracted to each other, so the more neutrons, [what?]

p^+ = proton

n^0 = neutron

e^- = electron

Stable nuclei most likely have even number of p^+ & even number of n^0 (Table 25-3)

Number of p^+	Number of n^0	Number of naturally occurring stable nuclides	Examples
even	even	163	^4He , ^{12}C , ^{16}O , ^{18}O , ^{20}Ne
even	odd	55	^3He , ^7Be , ^{13}C , ^{17}O , ^{21}Ne
odd	even	50	^7Li , ^{11}B , ^{14}N , ^{19}F , ^{23}Na
odd	odd	4	^2H , ^6Li , ^{10}B , ^{14}N

Shell model for nucleus gives magic numbers: 2, 8, 20, 28, 50, 82, 114 (for p^+), 126 (for n^0)

<http://www.nndc.bnl.gov/nudat2/>

Brookhaven National Laboratory puts # neutrons on x axis

§25-1 Nuclear reactions destroy atoms -- contrary to Dalton's Atomic Theory
But Mass number = # nucleons (p^+ & n^0) is unchanged

\implies mass is approximately conserved

Mass-energy is conserved. $\Delta E = (\Delta m)c^2$

Charge is conserved.

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D Kirk



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Table of Contents Chapter 25 Nuclear Chemistry Study Guide For Content Mastery

1. Understanding the eBook Chapter 25 Nuclear Chemistry Study Guide For Content Mastery
 - The Rise of Digital Reading Chapter 25 Nuclear Chemistry Study Guide For Content Mastery
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 25 Nuclear Chemistry Study Guide For Content Mastery
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 25 Nuclear Chemistry Study Guide For Content Mastery
 - User-Friendly Interface

4. Exploring eBook Recommendations from Chapter 25 Nuclear Chemistry Study Guide For Content Mastery
 - Personalized Recommendations
 - Chapter 25 Nuclear Chemistry Study Guide For Content Mastery User Reviews and Ratings
 - Chapter 25 Nuclear Chemistry Study Guide For Content Mastery and Bestseller Lists
5. Accessing Chapter 25 Nuclear Chemistry Study Guide For Content Mastery Free and Paid eBooks
 - Chapter 25 Nuclear Chemistry Study Guide For Content Mastery Public Domain eBooks
 - Chapter 25 Nuclear Chemistry Study Guide For Content Mastery eBook Subscription Services
 - Chapter 25 Nuclear Chemistry Study Guide For Content Mastery Budget-Friendly Options
6. Navigating Chapter 25 Nuclear Chemistry Study Guide For Content Mastery eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 25 Nuclear Chemistry Study Guide For Content Mastery Compatibility with Devices
 - Chapter 25 Nuclear Chemistry Study Guide For Content Mastery Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 25 Nuclear Chemistry Study Guide For Content Mastery
 - Highlighting and Note-Taking Chapter 25 Nuclear Chemistry Study Guide For Content Mastery
 - Interactive Elements Chapter 25 Nuclear Chemistry Study Guide For Content Mastery
8. Staying Engaged with Chapter 25 Nuclear Chemistry Study Guide For Content Mastery
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 25 Nuclear Chemistry Study Guide For Content Mastery
9. Balancing eBooks and Physical Books Chapter 25 Nuclear Chemistry Study Guide For Content Mastery
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 25 Nuclear Chemistry Study Guide For Content Mastery
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 25 Nuclear Chemistry Study Guide For Content Mastery
 - Setting Reading Goals Chapter 25 Nuclear Chemistry Study Guide For Content Mastery
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Chapter 25 Nuclear Chemistry Study Guide For Content Mastery
 - Fact-Checking eBook Content of Chapter 25 Nuclear Chemistry Study Guide For Content Mastery
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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