

## Section 12.1

### Section 12.1

Answer complete answer multiple-choice questions that are provided.

1. \_\_\_\_\_
  - a. \_\_\_\_\_
    - (a) correct, in terms of all the reactants and products.
    - (b) correct, in terms of moles of reactants and products.
    - (c) correct, in terms of work done on work done by the system.
    - (d) correct, in terms of enthalpy change,  $\Delta H$ .
  - b. \_\_\_\_\_
    - (a) the heat capacity of the system.
    - (b) the mass of the system.
    - (c) the number of moles of hydrogen that will react with a certain quantity of oxygen.
    - (d) the initial enthalpy of the system.
  - c. \_\_\_\_\_
    - (a) the number of moles of hydrogen that will react with a certain quantity of oxygen.
    - (b) the energy released in the reaction.
    - (c) the initial enthalpy of the system.
    - (d) the number of moles of hydrogen that will react with a certain quantity of oxygen.
  - d. \_\_\_\_\_
    - (a) the number of moles of hydrogen that will react with a certain quantity of oxygen.
    - (b) the initial enthalpy of the system.
    - (c) the number of moles of hydrogen that will react with a certain quantity of oxygen.
    - (d) the number of moles of hydrogen that will react with a certain quantity of oxygen.

2. Answer the two questions posed by the following balanced equation:



a. Balance the equation:



b. What is the coefficient of  $\text{H}_2$  in the balanced equation?

c. How many moles of  $\text{H}_2$  are required to produce 1.0 mole of  $\text{H}_2\text{O}$ ?

d. How many moles of  $\text{O}_2$  are required to produce 1.0 mole of  $\text{H}_2\text{O}$ ?

e. How many moles of  $\text{H}_2$  are required to produce 1.0 mole of  $\text{H}_2\text{O}$ ?

# Chapter 9 Review Stoichiometry Answers Section 1

**Terry C. Jones**



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