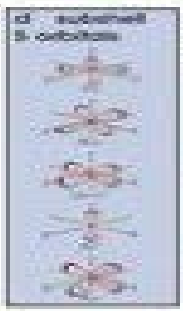


# CHEMISTRY

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## PERIODIC TABLE OF THE ELEMENTS



### ATOMIC STRUCTURE

Atomic Number,  $Z$ , of an element is the number of protons in the nucleus. For a neutral atom,  $Z = +$  the number of electrons. The mass number,  $A$ , is the sum of protons and neutrons in the nucleus. Isotopes of an element have the same  $Z$ , different  $A$ . The atomic mass is the weighted average of isotopes. The atomic weight is the weighted average of isotopes for a given element.

### ATOMIC QUANTUM NUMBERS AND ORBITALS

Principal quantum number,  $n$ , is the number of shells.  $n = 1, 2, 3, \dots$ . Each shell has a certain energy.  $n = 1$  is the lowest energy shell.  $n = 2$  is the second energy shell.  $n = 3$  is the third energy shell. Each shell has a certain number of orbitals.  $n = 1$  has 1 orbital.  $n = 2$  has 4 orbitals.  $n = 3$  has 9 orbitals.

### NUCLEAR CHEMISTRY

Nuclear reactions occur when particles (alpha, beta, gamma) are emitted from the nucleus. Alpha particles are  ${}^4_2\text{He}$ . Beta particles are  ${}^0_{-1}\text{e}$ . Gamma rays are  $\gamma$ . Radioactive decay is the process by which an unstable nucleus loses energy by emitting radiation. The half-life of a radioactive substance is the time it takes for half of the substance to decay.

### MANY-ELECTRON ATOMS

Orbitals and quantum numbers are used to describe the arrangement of electrons in an atom. The Aufbau principle states that electrons fill orbitals in order of increasing energy. The Pauli exclusion principle states that no two electrons can have the same set of quantum numbers. Hund's rule states that electrons fill degenerate orbitals singly before pairing up.

n	l	m <sub>l</sub>	m <sub>s</sub>	Orbital
1	0	0	$+\frac{1}{2}$	1s orbital
1	0	0	$-\frac{1}{2}$	1s orbital
2	0	0	$+\frac{1}{2}$	2s orbital
2	0	0	$-\frac{1}{2}$	2s orbital
2	1	1	$+\frac{1}{2}$	2p orbital
2	1	0	$+\frac{1}{2}$	2p orbital
2	1	-1	$+\frac{1}{2}$	2p orbital
2	1	1	$-\frac{1}{2}$	2p orbital
2	1	0	$-\frac{1}{2}$	2p orbital
2	1	-1	$-\frac{1}{2}$	2p orbital
3	0	0	$+\frac{1}{2}$	3s orbital
3	0	0	$-\frac{1}{2}$	3s orbital
3	1	1	$+\frac{1}{2}$	3p orbital
3	1	0	$+\frac{1}{2}$	3p orbital
3	1	-1	$+\frac{1}{2}$	3p orbital
3	1	1	$-\frac{1}{2}$	3p orbital
3	1	0	$-\frac{1}{2}$	3p orbital
3	1	-1	$-\frac{1}{2}$	3p orbital
3	2	2	$+\frac{1}{2}$	3d orbital
3	2	1	$+\frac{1}{2}$	3d orbital
3	2	0	$+\frac{1}{2}$	3d orbital
3	2	-1	$+\frac{1}{2}$	3d orbital
3	2	2	$-\frac{1}{2}$	3d orbital
3	2	1	$-\frac{1}{2}$	3d orbital
3	2	0	$-\frac{1}{2}$	3d orbital
3	2	-1	$-\frac{1}{2}$	3d orbital
3	2	2	$+\frac{1}{2}$	3d orbital
3	2	1	$+\frac{1}{2}$	3d orbital
3	2	0	$+\frac{1}{2}$	3d orbital
3	2	-1	$+\frac{1}{2}$	3d orbital
3	2	2	$-\frac{1}{2}$	3d orbital
3	2	1	$-\frac{1}{2}$	3d orbital
3	2	0	$-\frac{1}{2}$	3d orbital
3	2	-1	$-\frac{1}{2}$	3d orbital
3	2	2	$+\frac{1}{2}$	3d orbital
3	2	1	$+\frac{1}{2}$	3d orbital
3	2	0	$+\frac{1}{2}$	3d orbital
3	2	-1	$+\frac{1}{2}$	3d orbital
3	2	2	$-\frac{1}{2}$	3d orbital
3	2	1	$-\frac{1}{2}$	3d orbital
3	2	0	$-\frac{1}{2}$	3d orbital
3	2	-1	$-\frac{1}{2}$	3d orbital

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web in alpha decay an alpha particle is ejected from an unstable nucleus so here's our unstable nucleus uranium 238 an alpha particle has the same composition as a helium nucleus we saw the helium nucleus in the previous video there are two protons in the helium nucleus and two neutrons

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web sep 12 2022 the half life of the nuclei and thus the material is  $t_{1/2} = 2.693 \times 10^6$  million years 35 a the decay constant is  $\lambda = 1.99 \times 10^{-5} \text{ s}^{-1}$  b since strontium 91 has an atomic mass of 90.90 g

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web nuclear equations represent the reactants and products in radioactive decay nuclear fission or nuclear fusion instead of chemical equations where it shows the different number of elements is conserved in a reaction in a nuclear reaction the atomic mass and proton number are conserved