

Name Key

1. Hydrogen, H_2 , is used as rocket fuel. The hydrogen is burned in oxygen, O_2 , according to the following equation. What is the enthalpy change if 200.0 g of hydrogen, H_2 is burned?



$$\frac{200.0 \text{ g } H_2}{2.016 \text{ g } H_2} \times \frac{1 \text{ mol } H_2}{1} = 99.2 \text{ mol } H_2$$

$$99.2 \text{ mol } H_2 \times \frac{-484 \text{ kJ}}{2 \text{ mol } H_2} = -24,006 \text{ kJ}$$

-24,006 kJ

2. Carbon disulfide burns in air, producing carbon dioxide and sulfur dioxide.



What is ΔH for the following reaction?



3. Following is the equation for the combustion of acetone:



How much heat will be given off (in kcal) if 4.5 mol of acetone is burned completely?

$$\frac{-853.6 \text{ kcal}}{2 \text{ mol } C_3H_6O} \times 4.5 \text{ mol } C_3H_6O = -1920.6 \text{ kcal}$$

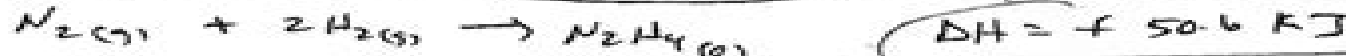
Exothermic

4. The specific heat of iron is $0.110 \text{ cal} / \text{g} \cdot ^\circ\text{C}$. How much heat must be added to 5.2 g of iron at 25°C to raise its temperature to 225°C ?

$$\text{heat} = q = m c \Delta T = (5.2 \text{ g}) \left(\frac{0.110 \text{ cal}}{\text{g} \cdot ^\circ\text{C}} \right) (225 - 25^\circ\text{C}) = 114.4 \text{ cal}$$

Endothermic

5. Hydrazine, N_2H_4 , is a colorless liquid used as a rocket fuel. What is the enthalpy change for the process in which hydrazine is formed from its elements (reaction (1) below)? Use Hess's Law to calculate ΔH for reaction (1) given the data below.



Chemistry Practice Problems With Answers

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Table of Contents Chemistry Practice Problems With Answers

1. Understanding the eBook Chemistry Practice Problems With Answers
 - The Rise of Digital Reading Chemistry Practice Problems With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Practice Problems With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Chemistry Practice Problems With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Practice Problems With Answers
 - Personalized Recommendations
 - Chemistry Practice Problems With Answers User Reviews and Ratings
 - Chemistry Practice Problems With Answers and Bestseller Lists

5. Accessing Chemistry Practice Problems With Answers Free and Paid eBooks
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 - Highlighting and Note-Taking Chemistry Practice Problems With Answers
 - Interactive Elements Chemistry Practice Problems With Answers
8. Staying Engaged with Chemistry Practice Problems With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Practice Problems With Answers
9. Balancing eBooks and Physical Books Chemistry Practice Problems With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Practice Problems With Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Practice Problems With Answers
 - Setting Reading Goals Chemistry Practice Problems With Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Practice Problems With Answers
 - Fact-Checking eBook Content of Chemistry Practice Problems With Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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