

# SEMESTER 1 FINAL REVIEW – CHEMISTRY

## 2015-16

Please ensure that you know how to answer these questions correctly, these are all good example problems for your final exam. Study your notes, labs, homework assignments and use the PowerPoint presentations that are online. Any of this information is fair game to be asked on the final exam. Good luck and ask questions in class during the review times if you have them. Don't wait to the last minute to start reviewing!

1. What is the number of significant figures in 1007000.0
2. What is the difference between accuracy and precision?
3. What is Avogadro's number? Why is it useful?
4. Find the molecular weight of  $\text{CO}_2$ ,  $\text{NH}_3$ ,  $\text{SF}_6$  ?
5. Be able to use density to calculate mass and or volume for various objects. Also, understand that density can be used at times with moles and particles through the  $D=m/v$  formula.
6. What is a mole? Be able to explain how the mole is a central connecting idea between mass, particles and gases @STP
7. What is the molar volume of a gas at STP?
8. What are the seven diatomic elements?
9. Define: atom, molecule, diatomic gas, monoatomic gas, element, compound, STP
10. What elements form the following bonds: ionic, covalent
11. Give an example of the following (both name and formula)
  - a. an ionic compound
  - b. a covalent compound
  - c. a metal
  - d. a nonmetal
  - e. a diatomic gas
12. Name the following:

a) NaOH	b) $\text{SO}_2$	c) $\text{C}_2\text{SO}_4$	d) HCl	e) $\text{Ca}(\text{NO}_3)_2$	f) $\text{NH}_4$
g) $\text{N}_2\text{O}_5$	h) $\text{Al}_2(\text{SO}_4)_3$	i) $\text{K}_3\text{PO}_4$	j) $\text{Na}_2\text{O}$	k) CO	l) $\text{CO}_2$
13. Write formulas for following
  - a) dinitrogen trioxide
  - b) potassium sulfate
  - c) hydrogen phosphate
  - d) sodium bromide
  - e) ammonium acetate
  - f) manganese IV nitrate
  - g) iron III hydroxide
  - h) aluminum carbonate
14. Write the following in correct scientific notation:

a) 476	b) 0.0000000002223	c) 3240000000000000
d) $462.9 \times 10^{23}$	e) $0.000000345 \times 10^{12}$	
f) $5677.9 \times 10^{-12}$	g) $0.000000223 \times 10^{-6}$	
15. How many oxygen atoms are there in 420 grams of carbon dioxide?
16. What is the mass of  $1.2 \times 10^{24}$  molecules of bromine?
17. What is the volume occupied by 243 grams of oxygen gas at STP?
18. How do you convert from:
  - a) moles  $\rightarrow$  grams
  - b) grams  $\rightarrow$  moles
  - c) moles  $\rightarrow$  molecules
  - d) moles  $\rightarrow$  liters of gas at STP
  - e) moles  $\rightarrow$  liters of solution

# Chemistry Semester 1 Review Answers

**Joacim Rocklöv**



## **Chemistry Semester 1 Review Answers:**

*The School Review*, 1916      *The Midland Druggist and Pharmaceutical Review*, 1911      **Midland Druggist and the Pharmaceutical Review**, 1911      *The Software Encyclopedia 2000* Bowker Editorial Staff, 2000-05      **Instructional Tests in Chemistry** Earl Rouse Glenn, Louis E. Welton, 1929      **Catholic Educational Review** Edward Aloysius Pace, Thomas Edward Shields, 1927      *5 Steps to a 5 AP Chemistry* John Moore, 2003-08-22 For the more than one million students taking the AP exams each year Boxed quotes offering advice from students who have aced the exams and from AP teachers and college professors Sample tests that closely simulate real exams Review material based on the contents of the most recent tests Icons highlighting important facts vocabulary and frequently asked questions Websites and links to valuable online test resources along with author e mail addresses for students with follow up questions Authors who are either AP course instructors or exam developers      **Educational Review**, 1927      *Fundamentals of Probability* Saeed Ghahramani, 2015-11-04 *Fundamentals of Probability with Stochastic Processes Third Edition* teaches probability in a natural way through interesting and instructive examples and exercises that motivate the theory definitions theorems and methodology The author takes a mathematically rigorous approach while closely adhering to the historical development of probability      **The World Review**, 1929      **Artificial Intelligence in Education. Posters and Late Breaking Results, Workshops and Tutorials, Industry and Innovation Tracks, Practitioners, Doctoral Consortium and Blue Sky** Andrew M. Olney, Irene-Angelica Chounta, Zitao Liu, Olga C. Santos, Ig Ibert Bittencourt, 2024-07-01 This volume constitutes poster papers and late breaking results presented during the 25th International Conference on Artificial Intelligence in Education AIED 2024 which took place in Recife Brazil during July 8 12 2024 The 18 full papers and 92 short papers were carefully reviewed and selected from 200 submissions They are organized in topical sections as follows Part One Blue Sky Industry Innovation and Practitioner Wide AIED and Late Breaking Results Part Two Late Breaking Results Doctoral Consortium Workshops and Tutorials      *Educational Review* Nicholas Murray Butler, Frank Pierrepont Graves, William McAndrew, 1927 Vols 19 34 include Bibliography of education for 1899 1906 compiled by James I Wyer and others      **Annual Catalogue** College of Hawaii, 1910      **The New SAFE Act** United States. Congress. Senate. Committee on Health, Education, Labor, and Pensions. Subcommittee on Employment, Safety, and Training, 1999 Distributed to some depository libraries in microfiche      *Science Tests and Reviews* Buros Center, 1975 *Science Tests and Reviews* consisting of science sections of the first seven MMYs and *Tests in Print II* includes 217 original test reviews written by 81 specialists 18 excerpted test reviews 270 references on the construction use and validity of specific tests a bibliography on in print science tests references for specific tests cumulative name indexes for specific tests with references a publishers directory title index name index and a scanning index The 97 tests covered fall into the following categories 23 general 14 biology 35 chemistry 3 geology 6 miscellaneous and 16 physics      *The Detroit Educational Bulletin* Detroit (Mich.). Board of Education, 1926 Vols 2

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may 3 2012 unsur gas mulia adalah unsur unsur yang terdapat pada golongan viii a sistem periodik yaitu helium he neon ne argon ar kripton kr ksenon xe dan radon rn kelompok ini disebut gas mulia karena sifatnya yang sukar bereaksi

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gas mulia ditemukan di dekat golongan halogen dalam tabel periodik karena unsur gas mulia memiliki konfigurasi elektron yang penuh unsur unsur tersebut tidak reaktif dan senyawanya tidak dikenal akibatnya gas gas ini dikenal dengan gas inert

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sifat sifat gas mulia wujud gas mulia unsur gas mulia terdapat sebagai gas tak berwarna yang monoatomik molekul yg terdiri dari 1 atom ini erat kaitannya dengan struktur elektron oktet dan duplet dari gas mulia

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dec 3 2015 gas mulia dan halogen merupakan unsur kimia yang sifatnya sangat bertolak belakang gas mulia dinamakan mulia karena seperti logam mulia sukar sekali bereaksi sementara halogen halos garam artinya adalah pembentuk garam halogen memiliki 6 bilangan oksidasi sehingga mudah sekali berreaksi dengan unsur lainnya gas mulia noble

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pembuatan gas mulia gas helium mempunyai titik didih yang sangat rendah yaitu 268 8 0c sehingga pemisahan gas helium dari gas alam dilakukan dengan cara pengembunan liquefaction sampai gas alam akan mencair sekitar 156 0c dan gas helium terpisah dari gas alam pembuatan ne ar kr dan xe dilakukan dengan proses destilasi bertingkat

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golongan gas mulia dan halogen beberapa golongan unsur memiliki nama khusus seperti golongan viiia gas mulia dan viia halogen golongan tersebut memiliki nama khusus seperti sifat priodik sifat fisik dan sifat kimia kekhas sifat ke 2 golongan tersebut dapat membantu kita unsur unsur yang lain dalam 1 golongan semakin ke atas dan dalam

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oct 7 2021 disebut gas mulia karena semua unsur pada golongan ini berwujud gas dan memiliki konfigurasi elektron yang sangat stabil sehingga akan sangat sulit untuk bereaksi dengan unsur lainnya unsur unsur gas mulia antara lain adalah helium he neon ne argon ar kripton kr xenon xe dan radon rn

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nov 26 2015 1 wujud gas mulia unsur gas mulia terdapat sebagai gas tak berwarna yang monoatomik ini erat kaitannya dengan struktur elektronoktet dan duplet dari gas mulia sedangkan wujud gas pada suhu kamar disebabkan titik

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