

Gravimetric Analysis Example – try it!

- An excess of aqueous sodium phosphate is added to a 50.0 mL sample of copper (II) chloride, with unknown concentration. A precipitate forms, as shown in the molecular equation below:



A student purifies, dries, and weighs the solid precipitate. She weighs the solid and finds she has collected 18.0 grams of $\text{Cu}_3(\text{PO}_4)_2$. Based on this data, determine the molarity of the CuCl_2 solution.

$$18.0 \text{ g } \cancel{\text{Cu}_3(\text{PO}_4)_2} \times \frac{1 \text{ mol } \cancel{\text{Cu}_3(\text{PO}_4)_2}}{380.4 \text{ g } \cancel{\text{Cu}_3(\text{PO}_4)_2}} \times \frac{3 \text{ mol } \text{CuCl}_2}{1 \cancel{\text{Cu}_3(\text{PO}_4)_2}} = 0.142 \text{ mol } \text{CuCl}_2$$

$$M = \frac{\text{mol}}{\text{L}} = \frac{0.142 \text{ mol}}{0.0500 \text{ L}} = 2.84$$

Stoichiometry And Gravimetric Analysis Lab Answers

N Noddings



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Working with Chemistry Donald J. Wink, Sharon Fetzer-Gislason, Julie Ellefson Kuehn, 2004-02-20 With this modular laboratory program students build skills using important chemical concepts and techniques to the point where they are able to design a solution to a scenario drawn from a professional environment The scenarios are drawn from the lives of people who work with chemistry every day ranging from field ecologists to chemical engineers and include many health professionals as well

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