

Name _____ Sex _____ Class _____

CHAPTER 12

STUDY GUIDE FOR CONTENT MASTERY

Section 12.3 Limiting Reactants

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then use data in the table. Show your calculations for questions 15–17 in the space below the table.



The molar masses of O_2 , H_2 , and H_2O are 32.00 g/mol, 2.02 g/mol, and 18.02 g/mol, respectively.

Amount of O_2	Amount of H_2	Amount of H_2O	Limiting Reactant	Amount of Excess Reactant
1 molecule	1 molecule	1 molecule	none	none
2 molecules	2 molecules	2 molecules	H_2	1 molecule O_2
2 molecules	3 molecules	3 molecules	O_2	1 molecule H_2
1.00 mol	2.00 mol	2.00 mol	H_2	1.00 mol O_2
2.00 mol	4.00 mol	4.00 mol	H_2	2.00 mol O_2
3.00 mol	6.00 mol	6.00 mol	H_2	3.00 mol O_2
4.00 mol	8.00 mol	8.00 mol	H_2	4.00 mol O_2
5.00 mol	10.00 mol	10.00 mol	H_2	5.00 mol O_2
6.00 mol	12.00 mol	12.00 mol	H_2	6.00 mol O_2
7.00 mol	14.00 mol	14.00 mol	H_2	7.00 mol O_2
8.00 mol	16.00 mol	16.00 mol	H_2	8.00 mol O_2
9.00 mol	18.00 mol	18.00 mol	H_2	9.00 mol O_2
10.00 mol	20.00 mol	20.00 mol	H_2	10.00 mol O_2
11.00 mol	22.00 mol	22.00 mol	H_2	11.00 mol O_2
12.00 mol	24.00 mol	24.00 mol	H_2	12.00 mol O_2
13.00 mol	26.00 mol	26.00 mol	H_2	13.00 mol O_2
14.00 mol	28.00 mol	28.00 mol	H_2	14.00 mol O_2
15.00 mol	30.00 mol	30.00 mol	H_2	15.00 mol O_2

balanced equation mole ratio = 2 mol H_2O /1 mol O_2

16.00 g O_2 = 1 mol O_2 /32.00 g O_2 = 0.500 mol O_2

20.00 g H_2 = 1 mol H_2 /2.02 g H_2 = 9.90 mol H_2

actual mole ratio = 9.90 mol H_2 /0.500 mol O_2 = 19.8 mol H_2 /1.00 mol O_2

Because the actual mole ratio of H_2 / O_2 is larger than the balanced equation mole ratio of H_2 / O_2 , there is an excess of H_2 ; O_2 is the limiting reactant.

Mass of H_2O used = 0.500 mol H_2O = 2 mol H_2O /1 mol O_2 = 0.500 mol H_2O

0.500 mol H_2O = 9.01 g H_2O /1 mol H_2O = 9.01 g H_2O

Mass of O_2 produced = 0.500 mol O_2 = 16.00 g O_2 /1 mol O_2 = 16.00 g O_2

Excess H_2 = 20.00 g H_2 - 9.01 g H_2 = 10.99 g H_2

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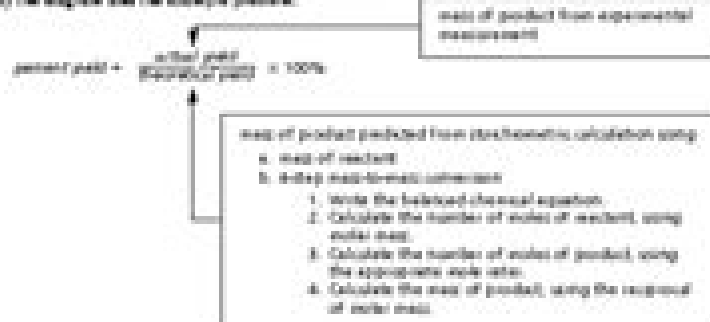
CHAPTER 12

STUDY GUIDE FOR CONTENT MASTERY

Section 12.4 Percent Yield

In your textbook, read about the yield of products.

Study the diagram and the example problem.



Example Problem: The following chemical equation represents the production of gallium nitride, a substance used in the manufacturing of some semiconductors for lasers.



In one experiment, the reaction yielded 7.43 g of the nitride from a 1.00-g sample of gallium. Determine the percent yield of this reaction. The molar masses of Ga and GaN are 69.72 g/mol and 122.62 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- percent yield solution $\frac{7.43 \text{ g } GaN}{1.00 \text{ g } Ga} \times 100\%$
- mass of reactant 1.00 g Ga $\frac{69.72 \text{ g Ga}}{1 \text{ mol Ga}} \times 1.00 \text{ g Ga} = 1 \text{ mol Ga}$
- number of moles of reactant 1.00 g Ga = $\frac{69.72 \text{ g Ga}}{1 \text{ mol Ga}} \times 1.00 \text{ g Ga} = 1 \text{ mol Ga}$
- number of moles of product 0.667 mol Ga = $\frac{2 \text{ mol GaN}}{3 \text{ mol Ga}} \times 1 \text{ mol Ga} = 0.667 \text{ mol GaN}$
- theoretical yield 0.667 mol GaN = $\frac{122.62 \text{ g GaN}}{1 \text{ mol GaN}} \times 0.667 \text{ mol GaN} = 84.0 \text{ g GaN}$
- percent yield $\frac{7.43 \text{ g GaN}}{84.0 \text{ g GaN}} \times 100\% = 8.84\%$

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VM Jensen



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Table of Contents Study Guide For Content Mastery Answer Key Chapter 12

1. Understanding the eBook Study Guide For Content Mastery Answer Key Chapter 12
 - The Rise of Digital Reading Study Guide For Content Mastery Answer Key Chapter 12
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide For Content Mastery Answer Key Chapter 12
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide For Content Mastery Answer Key Chapter 12
 - User-Friendly Interface

4. Exploring eBook Recommendations from Study Guide For Content Mastery Answer Key Chapter 12
 - Personalized Recommendations
 - Study Guide For Content Mastery Answer Key Chapter 12 User Reviews and Ratings
 - Study Guide For Content Mastery Answer Key Chapter 12 and Bestseller Lists
5. Accessing Study Guide For Content Mastery Answer Key Chapter 12 Free and Paid eBooks
 - Study Guide For Content Mastery Answer Key Chapter 12 Public Domain eBooks
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 - Study Guide For Content Mastery Answer Key Chapter 12 Budget-Friendly Options
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 - Study Guide For Content Mastery Answer Key Chapter 12 Enhanced eBook Features
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 - Highlighting and Note-Taking Study Guide For Content Mastery Answer Key Chapter 12
 - Interactive Elements Study Guide For Content Mastery Answer Key Chapter 12
8. Staying Engaged with Study Guide For Content Mastery Answer Key Chapter 12
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide For Content Mastery Answer Key Chapter 12
9. Balancing eBooks and Physical Books Study Guide For Content Mastery Answer Key Chapter 12
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide For Content Mastery Answer Key Chapter 12
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Study Guide For Content Mastery Answer Key Chapter 12
 - Setting Reading Goals Study Guide For Content Mastery Answer Key Chapter 12
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Study Guide For Content Mastery Answer Key Chapter 12
 - Fact-Checking eBook Content of Study Guide For Content Mastery Answer Key Chapter 12
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
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