

Balancing Chemical Equations Worksheet

1. $\underline{2} \text{ H}_2 + \underline{\quad} \text{ O}_2 \rightarrow \underline{2} \text{ H}_2\text{O}$
2. $\underline{\quad} \text{ N}_2 + \underline{3} \text{ H}_2 \rightarrow \underline{2} \text{ NH}_3$
3. $\underline{\quad} \text{ S}_8 + \underline{12} \text{ O}_2 \rightarrow \underline{8} \text{ SO}_3$
4. $\underline{2} \text{ N}_2 + \underline{\quad} \text{ O}_2 \rightarrow \underline{2} \text{ N}_2\text{O}$
5. $\underline{2} \text{ HgO} \rightarrow \underline{2} \text{ Hg} + \underline{\quad} \text{ O}_2$
6. $\underline{4} \text{ CO}_2 + \underline{6} \text{ H}_2\text{O} \rightarrow \underline{\quad} \text{ C}_6\text{H}_{12}\text{O}_6 + \underline{6} \text{ O}_2$
7. $\underline{\quad} \text{ Zn} + \underline{2} \text{ HCl} \rightarrow \underline{\quad} \text{ ZnCl}_2 + \underline{\quad} \text{ H}_2$
8. $\underline{\quad} \text{ SiCl}_4 + \underline{4} \text{ H}_2\text{O} \rightarrow \underline{\quad} \text{ H}_4\text{SiO}_4 + \underline{4} \text{ HCl}$
9. $\underline{2} \text{ Na} + \underline{2} \text{ H}_2\text{O} \rightarrow \underline{2} \text{ NaOH} + \underline{\quad} \text{ H}_2$
10. $\underline{2} \text{ H}_3\text{PO}_4 \rightarrow \underline{\quad} \text{ H}_4\text{P}_2\text{O}_7 + \underline{\quad} \text{ H}_2\text{O}$
11. $\underline{\quad} \text{ C}_{10}\text{H}_{16} + \underline{8} \text{ Cl}_2 \rightarrow \underline{10} \text{ C} + \underline{16} \text{ HCl}$
12. $\underline{\quad} \text{ CO}_2 + \underline{2} \text{ NH}_3 \rightarrow \underline{\quad} \text{ OC}(\text{NH}_2)_2 + \underline{\quad} \text{ H}_2\text{O}$
13. $\underline{4} \text{ Si}_2\text{H}_3 + \underline{17} \text{ O}_2 \rightarrow \underline{8} \text{ SiO}_2 + \underline{6} \text{ H}_2\text{O}_3$
14. $\underline{2} \text{ Al}(\text{OH})_3 + \underline{3} \text{ H}_2\text{SO}_4 \rightarrow \underline{\quad} \text{ Al}_2(\text{SO}_4)_3 + \underline{6} \text{ H}_2\text{O}$
15. $\underline{4} \text{ Fe} + \underline{3} \text{ O}_2 \rightarrow \underline{2} \text{ Fe}_2\text{O}_3$
16. $\underline{\quad} \text{ Fe}_2(\text{SO}_4)_3 + \underline{6} \text{ KOH} \rightarrow \underline{3} \text{ K}_2\text{SO}_4 + \underline{2} \text{ Fe}(\text{OH})_3$
17. $\underline{2} \text{ C}_7\text{H}_6\text{O}_2 + \underline{15} \text{ O}_2 \rightarrow \underline{14} \text{ CO}_2 + \underline{6} \text{ H}_2\text{O}$
18. $\underline{\quad} \text{ H}_2\text{SO}_4 + \underline{8} \text{ HI} \rightarrow \underline{\quad} \text{ H}_2\text{S} + \underline{4} \text{ I}_2 + \underline{4} \text{ H}_2\text{O}$
19. $\underline{4} \text{ FeS}_2 + \underline{11} \text{ O}_2 \rightarrow \underline{2} \text{ Fe}_2\text{O}_3 + \underline{8} \text{ SO}_2$
20. $\underline{2} \text{ Al} + \underline{3} \text{ FeO} \rightarrow \underline{\quad} \text{ Al}_2\text{O}_3 + \underline{3} \text{ Fe}$
21. $\underline{\quad} \text{ Fe}_2\text{O}_3 + \underline{3} \text{ H}_2 \rightarrow \underline{2} \text{ Fe} + \underline{3} \text{ H}_2\text{O}$
22. $\underline{\quad} \text{ Na}_2\text{CO}_3 + \underline{2} \text{ HCl} \rightarrow \underline{2} \text{ NaCl} + \underline{\quad} \text{ H}_2\text{O} + \underline{\quad} \text{ CO}_2$
23. $\underline{2} \text{ K} + \underline{\quad} \text{ Br}_2 \rightarrow \underline{2} \text{ KBr}$
24. $\underline{\quad} \text{ C}_7\text{H}_{16} + \underline{11} \text{ O}_2 \rightarrow \underline{7} \text{ CO}_2 + \underline{8} \text{ H}_2\text{O}$
25. $\underline{2} \text{ P}_4 + \underline{5} \text{ O}_2 \rightarrow \underline{2} \text{ P}_2\text{O}_5$

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Balancing Chemical Equations Answer Sheet:

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