

NAME _____ DATE _____ CLASS _____

CHAPTER 12 **STUDY GUIDE FOR CONTENT MASTERY**

Section 12.3 Limiting Reactants

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then complete the table. Show your calculations for questions 25–27 in the space below the table.



The molar masses of O_2 , NO , and NO_2 are 32.00 g/mol, 30.01 g/mol, and 46.01 g/mol, respectively.

Amount of O_2	Amount of NO	Amount of NO_2	Limiting Reactant	Amount and Mass of Excess Reactant
1 molecule	2 molecules	2 molecules	none	none
2 molecules	4 molecules	4 molecules	NO	2 molecules O_2
2 molecules	4 molecules	4 molecules	1. O_2	2. 4 molecules NO
1.00 mol	2.00 mol	3. 2.00 mol	4. none	5. none
4.00 mol	6.00 mol	7. 4.00 mol	8. NO	9. 2.00 mol O_2
0.00 mol	0.00 mol	10. 7.00 mol	11. NO	12. 1.50 mol O_2
1.00 mol	0.00 mol	13. 2.00 mol	14. O_2	15. 2.00 mol NO
0.200 mol	0.200 mol	16. 0.200 mol	17. NO	18. 0.400 mol O_2
32.00 g	60.02 g	19. 92.02 g	20. none	21. none
64.00 g	60.02 g	22. 66.02 g	23. O_2	24. 30.12 g NO
64.00 g	30.01 g	25. 28.76 g	26. O_2	27. 1.24 g NO

balanced equation mole ratio = 2 mol NO /1 mol O_2

16.00 g O_2 = 1 mol O_2 /32.00 g O_2 = 0.5125 mol O_2

20.00 g NO = 1 mol NO /30.01 g NO = 0.6664 mol NO

actual mole ratio = 0.6664 mol NO /0.5125 mol O_2 = 2.132 mol NO /1.000 mol O_2

Because the actual mole ratio of NO / O_2 is larger than the balanced equation mole ratio of NO / O_2 , there is an excess of NO ; O_2 is the limiting reactant.

Mass of NO used = 0.5125 mol O_2 = 2 mol NO /1 mol O_2 = 0.4250 mol NO

0.4250 mol NO = 30.01 g NO /1 mol NO = 12.76 g NO

Mass of NO_2 produced = 0.4250 mol NO_2 = 46.01 g NO_2 /1 mol NO_2 = 19.76 g NO_2

Excess NO = 20.00 g NO = 12.76 g NO = 7.24 g NO

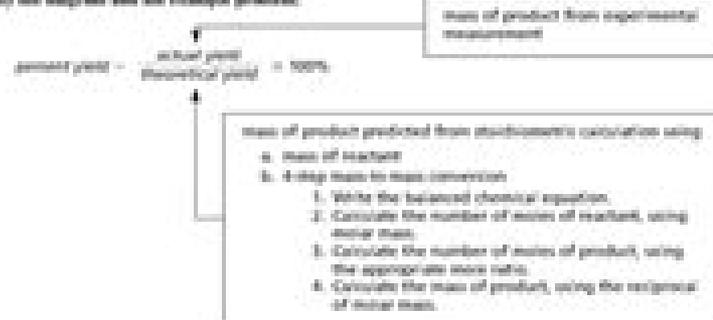
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CHAPTER 12 **STUDY GUIDE FOR CONTENT MASTERY**

Section 12.4 Percent Yield

In your textbook, read about the percent of products.

Study the diagram and the example problem.



Example Problem: The following chemical equation represents the production of yellow oxide, a substance used in the manufacturing of some semiconductor devices.



In one experiment, the reaction yielded 7.42 g of the oxide from a 7.00-g sample of gallium. Determine the percent yield of this reaction. The molar masses of Ga and Ga_2O_3 are 69.72 g/mol and 177.44 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- actual yield: unknown 7.42 g Ga_2O_3
- mass of reactant: 7.00 g Ga correct
- number of moles of reactant: 7.00 g Ga = $\frac{69.72 \text{ g } Ga}{1 \text{ mol } Ga}$ 7.00 g Ga = 1 mol Ga /69.72 g Ga
- number of moles of product: 0.100 mol Ga = $\frac{2 \text{ mol } Ga_2O_3}{1 \text{ mol } Ga}$ 0.100 mol Ga = 2 mol Ga_2O_3 /4 mol Ga
- theoretical yield: 0.200 mol Ga_2O_3 = $\frac{177.44 \text{ g } Ga_2O_3}{1 \text{ mol } Ga_2O_3}$ correct
- percent yield: $\frac{7.42 \text{ g } Ga_2O_3}{177.44 \text{ g } Ga_2O_3} \times 100$ 7.42 g Ga_2O_3 /177.44 g Ga_2O_3 = 100

Standardized Test Practice Chapter 12 Chemistry Answers

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GED Science For Dummies Murray Shukyn, Achim K. Krull, 2015-09-23 Passing the GED Science Test has never been easier Does the thought of taking the GED Science Test make you sweat Fear not With the help of GED Science Test For Dummies you'll get up to speed on the new structure and computer based format of the GED and gain the confidence and know how to pass the Science Test like a pro Packed with helpful guidance and instruction this hands on test prep guide covers the concepts covered on the GED Science Test and gives you ample practice opportunities to assess your understanding of Life Science Physical Science and Earth and Space Science Designed to test your understanding of the fundamentals of science reasoning and the ability to apply those fundamentals in realistic situations the GED Science Test can be tough for the uninitiated Luckily this fun and accessible guide breaks down each section of the exam into easily digestible parts making everything you'll encounter on exam day feel like a breeze Inside you'll find methods to sharpen your science vocabulary and data analysis skills tips on how to approach GED Science Test question types and formats practice questions and study exercises and a full length practice test to help you pinpoint where you need more study help Presents reviews of the GED Science test question types and basic computer skills Offers practice questions to assess your knowledge of each subject area Includes one full length GED Science practice test Provides scoring guidelines and detailed answer explanations Even if science is something that's always made you squeamish GED Science Test For Dummies makes it easy to pass this crucial exam and obtain your hard earned graduate equivalency diploma

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Study Guide and Solutions Manual to accompany General Organic and Biological Chemistry, 1e Kenneth W. Raymond,2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond s text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry **Using**

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