

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

## CHAPTER 12

## STUDY GUIDE FOR CONTENT MASTERY

## Section 12.3 Limiting Reactants

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then use data for table. Show your calculations for questions 15–17 in the space below the table.



The molar masses of  $O_2$ ,  $NO$ , and  $NO_2$  are 32.00 g/mol, 30.01 g/mol, and 46.01 g/mol, respectively.

Amount of $O_2$	Amount of $NO$	Amount of $NO_2$	Limiting Reactant	Amount of Excess Reactant
1 molecule	1 molecule	1 molecule	none	none
2 molecules	2 molecules	2 molecules	$NO$	1 molecule $O_2$
2 molecules	3 molecules	4 molecules	$O_2$	1 molecule $NO$
1.00 mol	2.00 mol	2.00 mol	$NO$	1.00 mol $O_2$
2.00 mol	4.00 mol	4.00 mol	$NO$	2.00 mol $O_2$
3.00 mol	6.00 mol	6.00 mol	$NO$	3.00 mol $O_2$
4.00 mol	8.00 mol	8.00 mol	$NO$	4.00 mol $O_2$
5.00 mol	10.00 mol	10.00 mol	$NO$	5.00 mol $O_2$
6.00 mol	12.00 mol	12.00 mol	$NO$	6.00 mol $O_2$
7.00 mol	14.00 mol	14.00 mol	$NO$	7.00 mol $O_2$
8.00 mol	16.00 mol	16.00 mol	$NO$	8.00 mol $O_2$
9.00 mol	18.00 mol	18.00 mol	$NO$	9.00 mol $O_2$
10.00 mol	20.00 mol	20.00 mol	$NO$	10.00 mol $O_2$
11.00 mol	22.00 mol	22.00 mol	$NO$	11.00 mol $O_2$
12.00 mol	24.00 mol	24.00 mol	$NO$	12.00 mol $O_2$
13.00 mol	26.00 mol	26.00 mol	$NO$	13.00 mol $O_2$
14.00 mol	28.00 mol	28.00 mol	$NO$	14.00 mol $O_2$
15.00 mol	30.00 mol	30.00 mol	$NO$	15.00 mol $O_2$
16.00 mol	32.00 mol	32.00 mol	$NO$	16.00 mol $O_2$
17.00 mol	34.00 mol	34.00 mol	$NO$	17.00 mol $O_2$
18.00 mol	36.00 mol	36.00 mol	$NO$	18.00 mol $O_2$
19.00 mol	38.00 mol	38.00 mol	$NO$	19.00 mol $O_2$
20.00 mol	40.00 mol	40.00 mol	$NO$	20.00 mol $O_2$

balanced equation mole ratio = 2 mol  $NO$ /1 mol  $O_2$

32.00 g  $O_2$  = 1 mol  $O_2$ /32.00 g  $O_2$  = 0.3125 mol  $O_2$

20.00 g  $NO$  = 1 mol  $NO$ /30.01 g  $NO$  = 0.6664 mol  $NO$

actual mole ratio = 0.6664 mol  $NO$ /0.3125 mol  $O_2$  = 2.132 mol  $NO$ /1.000 mol  $O_2$

Because the actual mole ratio of  $NO$ / $O_2$  is larger than the balanced equation mole ratio of  $NO$ / $O_2$ , there is an excess of  $NO$ / $O_2$  is the limiting reactant.

Mass of  $NO$  used = 0.3125 mol  $O_2$  = 2 mol  $NO$ /1 mol  $O_2$  = 0.6250 mol  $NO$

0.6250 mol  $NO$  = 30.01 g  $NO$ /1 mol  $NO$  = 18.75 g  $NO$

Mass of  $NO_2$  produced = 0.6250 mol  $NO_2$  = 46.01 g  $NO_2$ /1 mol  $NO_2$  = 28.75 g  $NO_2$

Excess  $NO$  = 20.00 g  $NO$  - 18.75 g  $NO$  = 1.24 g  $NO$

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

## CHAPTER 12

## STUDY GUIDE FOR CONTENT MASTERY

## Section 12.4 Percent Yield

In your textbook, read about the yield of products.

Study the diagram and the example problem.



**Example Problem:** The following chemical equation represents the production of gallium nitride, a substance used in the manufacturing of some semiconductors for lasers.



In one experiment, the reaction yielded 7.42 g of the oxide from a 1.00-g sample of gallium. Determine the percent yield of this reaction. The molar masses of Ga and  $GaN_3$  are 69.72 g/mol and 127.63 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

1. actual yield volume = 7.42 g  $GaN_3$
2. mass of reactant = 1.00 g Ga, 100% Ga
3. number of moles of reactant = 1.00 g Ga  $\times$   $\frac{69.72 \text{ g Ga}}{1 \text{ mol Ga}}$  = 7.00 g Ga = 1 mol Ga/69.72 g Ga
4. number of moles of product = 0.00 mol Ga  $\times$   $\frac{2 \text{ mol } GaN_3}{1 \text{ mol Ga}}$  =  $\frac{0.100 \text{ mol Ga} \times 2}{2 \text{ mol } Ga} = 0.100 \text{ mol Ga}$
5. theoretical yield = 0.0000 mol  $GaN_3$   $\times$   $\frac{127.63 \text{ g } GaN_3}{1 \text{ mol } GaN_3}$  = correct
6. percent yield =  $\frac{7.42 \text{ g } GaN_3}{127.63 \text{ g } GaN_3} \times 100$  =  $\frac{7.42 \text{ g } GaN_3}{67 \text{ g } GaN_3} \times 100$

# Study Guide For Content Mastery Answer Key Chapter 12

**Sabine Zange**



**Study Guide For Content Mastery Answer Key Chapter 12:**

This book delves into Study Guide For Content Mastery Answer Key Chapter 12. Study Guide For Content Mastery Answer Key Chapter 12 is an essential topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Study Guide For Content Mastery Answer Key Chapter 12, encompassing both the fundamentals and more intricate discussions.

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## **Study Guide For Content Mastery Answer Key Chapter 12**

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web annecy is part of four cantons annecy 1 annecy 2 annecy 3 and annecy 4 it is the prefecture of haute savoie since 2017 annecy has consisted of six delegate cities annecy annecy le vieux cran gevrier meythet pringy and seynod local government consists of a city council with 202 members the number of members in each commune