

**KEY**

3. The pH in a 0.25 M solution of the acid HBrO is 4.65. Using this, calculate the value of Ka for the acid HBrO. (4 marks)  $pH = 4.65$  so  $[H_3O^+] = \text{antilog}(-4.65) = 2.239 \times 10^{-5} M$



|     |                               |  |                         |                         |
|-----|-------------------------------|--|-------------------------|-------------------------|
| [I] | 0.25                          |  | 0                       | 0                       |
| [C] | $-2.239 \times 10^{-5}$       |  | $+2.239 \times 10^{-5}$ | $+2.239 \times 10^{-5}$ |
| [E] | $0.25 - 2.239 \times 10^{-5}$ |  | $2.239 \times 10^{-5}$  | $2.239 \times 10^{-5}$  |

4

$$K_a = \frac{[H_3O^+][BrO^-]}{[HBrO]}$$

$$= \frac{(2.239 \times 10^{-5})^2}{(0.25 - 2.239 \times 10^{-5})} = 2.0 \times 10^{-9}$$

Answer  $K_a = 2.0 \times 10^{-9}$

4. The pH in a solution of benzoic acid is 2.355. Determine the molar concentration of the benzoic acid. (4 marks)  $pH = 2.355$   $[H_3O^+] = \text{antilog}(-2.355) = 0.004416 M$



|     |                  |  |             |             |
|-----|------------------|--|-------------|-------------|
| [I] | $C_0$            |  | 0           | 0           |
| [C] | $-0.004416$      |  | $+0.004416$ | $+0.004416$ |
| [E] | $C_0 - 0.004416$ |  | $0.004416$  | $0.004416$  |

4

$$K_a = \frac{[H_3O^+][C_6H_5COO^-]}{[C_6H_5COOH]}$$

$$K_a = \frac{(0.004416)^2}{(C_0 - 0.004416)}$$

$$C_0 - 0.004416 = \frac{(0.004416)^2}{6.5 \times 10^{-5}} = 0.299976$$

$$C_0 = 0.299976 + 0.004416 = 0.30439 M$$

Answer  $C_0 = 0.30 M$

5. Find the value of Kb for the oxalate ion ( $C_2O_4^{2-}$ ). (1 mark)

$$K_b(C_2O_4^{2-}) = \frac{K_w}{K_a(HC_2O_4^-)} = \frac{1.00 \times 10^{-14}}{6.4 \times 10^{-5}} = 1.56 \times 10^{-10}$$

1

Answer  $K_b = 1.6 \times 10^{-10}$

6. The value of Kb for the weak base methylamine ( $CH_3NH_2$ ) is  $4.4 \times 10^{-4}$ . Calculate the value of Ka for the acid  $CH_3NH_3^+$ . (1 mark)

$$K_a(CH_3NH_3^+) = \frac{K_w}{K_b(CH_3NH_2)} = \frac{1.00 \times 10^{-14}}{4.4 \times 10^{-4}} = 2.27 \times 10^{-11}$$

1

Answer  $K_a = 2.3 \times 10^{-11}$

10

# Basic Stoichiometry Phet Lab Answer Key

**M Carnoy**



**Basic Stoichiometry Phet Lab Answer Key :**

## Whispering the Secrets of Language: An Mental Quest through **Basic Stoichiometry Phet Lab Answer Key**

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