

Review for Ch. 8 Test

8.1 Describing Chemical Reactions

- know how to write word equations (i.e. you have be able to name compounds, including acids, ionic, and covalent compounds from Ch. 7)
- know how to write a formula equation
- What 2 things do coefficients represent in an equation? **Relative number of moles of a substance ; relative number of molecules needed for a reaction to occur**
- know common symbols used in a reaction
- know 4 evidences of a chemical reaction:
 1. **energy in the form of heat or light**
 2. **production of a gas**
 3. **change in color**
 4. **formation of a precipitate**
- know the Law of conservation of matter; **matter can neither be created nor destroyed in a chemical reaction (which means, mass of reactants must equal the mass of products)**
- know what a chemical equation represents: **relative amounts of reactants and products**
- know how to balance chemical equations (combustion reactions are particularly tricky, practice!)

1. Complete the equation and write a word equation for the following reaction: $\text{ZnO} + \text{FeSO}_4 \rightleftharpoons$



Zinc oxide reacts with iron (II) sulfate to yield zinc sulfate and iron (II) oxide

2. Write a formula equation for the reaction of a solution of AgCl with potassium metal:



3. What elements are always diatomic gases at room temperature?

All of the halogens (Group 17) , H_2 , N_2 , O_2

4. Can you balance these reactions:



Describing Chemical Reactions Section Review Answer Key

Scott C. Dulebohn



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Table of Contents Describing Chemical Reactions Section Review Answer Key

1. Understanding the eBook Describing Chemical Reactions Section Review Answer Key
 - The Rise of Digital Reading Describing Chemical Reactions Section Review Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Describing Chemical Reactions Section Review Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Describing Chemical Reactions Section Review Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Describing Chemical Reactions Section Review Answer Key
 - Personalized Recommendations
 - Describing Chemical Reactions Section Review Answer Key User Reviews and Ratings
 - Describing Chemical Reactions Section Review Answer Key and Bestseller Lists
5. Accessing Describing Chemical Reactions Section Review Answer Key Free and Paid eBooks
 - Describing Chemical Reactions Section Review Answer Key Public Domain eBooks
 - Describing Chemical Reactions Section Review Answer Key eBook Subscription Services
 - Describing Chemical Reactions Section Review Answer Key Budget-Friendly Options

6. Navigating Describing Chemical Reactions Section Review Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Describing Chemical Reactions Section Review Answer Key Compatibility with Devices
 - Describing Chemical Reactions Section Review Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Describing Chemical Reactions Section Review Answer Key
 - Highlighting and Note-Taking Describing Chemical Reactions Section Review Answer Key
 - Interactive Elements Describing Chemical Reactions Section Review Answer Key
8. Staying Engaged with Describing Chemical Reactions Section Review Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Describing Chemical Reactions Section Review Answer Key
9. Balancing eBooks and Physical Books Describing Chemical Reactions Section Review Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Describing Chemical Reactions Section Review Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Describing Chemical Reactions Section Review Answer Key
 - Setting Reading Goals Describing Chemical Reactions Section Review Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Describing Chemical Reactions Section Review Answer Key
 - Fact-Checking eBook Content of Describing Chemical Reactions Section Review Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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