

Name: _____

Biology, Period _____

Midterm Review for Biology - 520

Chapter #1 - The Science of Biology (19 terms)

1.1 What Is Science?

Describe the steps used in scientific method.

- Hypothesis – If...then...
- Identify independent variable versus dependent variable
- Identify controlled variables
- Identify control group versus experimental group

1.2 Science in Context

Explain what a scientific theory is.

1.3 Studying Life

Identify the 8 characteristics of living things.

Convert within the metric system (cm – m - km; mL – L, g – kg)

Chapter #2 -The Chemistry of Life (39 terms)

2.1 The Nature of Matter

Identify the three subatomic particles found in atoms, and their location.

- Proton
- Neutron
- Electron

Explain how all of the isotopes of an element are similar and how they are different.

Define compound.

Describe the two main types of chemical bonds.

- Ionic
- Covalent

2.2 Properties of Water

Draw a water molecule showing the positive and negative sides.

Discuss the unique properties of water.

- Hydrogen bonding
- Adhesion
- Cohesion

Differentiate between solutions and suspensions.

Explain what acidic solutions and basic solutions are.

- pH

2.3 - Carbon Compounds

Describe the unique qualities of carbon.

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L Darling-Hammond



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Introduction to Biology Alan Axelrod, 1999 A complete guide with questions answers and practice tests in the field of biology

Modern Biology Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2006-01-01 **Modern Biology**, 2002

Study Guide Essential Biology with Physiology Edward J. Zalisko, 2003-07 Students can master key concepts and earn a better grade with the thought provoking exercises found in this study guide Study advice tables quizzes and crossword puzzles help students test their understanding of biology The Study Guide also includes references to student media activities on the Essential Biology CD ROM and Website

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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) + \text{C}_2\text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_4(\text{g}) + 2 \text{ H}_2(\text{g}) + \text{C}_2\text{H}_6(\text{g})$ $\Delta H = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\Delta H = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{g})$ $\Delta H = 283 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g})$

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web chemistry 120 hess s law worksheet calculate ΔH for the reaction $\text{C}_2\text{H}_4(\text{g}) + \text{H}_2(\text{g}) + \text{C}_2\text{H}_6(\text{g})$ from the following data
 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l})$ $\Delta H = 1411 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + 7 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l})$ $\Delta H = 1560 \text{ kJ}$
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\Delta H = 285 \text{ kJ}$
mole

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web thus for reaction 9 2 in which 1 mole of h_3o^+ and 1 mole of oh^- combine to form 2 moles of h_2o h 55.8 kj as shown note that the enthalpy change $h_{products} - h_{reactants}$ is positive if heat is absorbed that is if $h_{products}$

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