

9. Evaluate the experiments that led to the conclusion that electrons are negatively charged particles found in all matter.

The deflection toward positively charged plates demonstrated the negatively charged nature of electrons; the fact that changing the type of electrode or the type of gas used in the cathode ray tube did not affect the ray produced led to the conclusion that electrons are present in all matter.

10. Compare the relative charge and mass of each of the subatomic particles.

Particle	Relative charge	Relative mass
Electron	-1	1/1840
Proton	+1	1
Neutron	0	-1

11. **Calculate** What is the difference expressed in kilograms between the mass of a proton and the mass of an electron?

$$1.673 \times 10^{-24} \text{ g} - 9.11 \times 10^{-28} \text{ g} = 1.672 \times 10^{-24} \text{ g} \times 1 \text{ kg}/1000 \text{ g} = 1.672 \times 10^{-27} \text{ kg}$$

Section 4.3 How Atoms Differ

pages 115–121

Practice Problems

pages 116–121

12. How many protons and electrons are in each atom?
- radon, 86 protons and 86 electrons
 - magnesium, 12 protons and 12 electrons
13. An atom of an element contains 66 electrons. Which element is it?
dysprosium
14. An atom of an element contains 14 protons. Which element is it?
silicon

15. **Challenge** Do the atoms shown in the figure to the right have the same atomic number?



Yes, 9

16. Determine the number of protons, electrons, and neutrons for isotopes b through f, in the table above. Name each isotope, and write its symbol.

	Protons and Electrons	Neutrons	Isotope	Symbol
b.	20	26	calcium-46	$^{46}_{20}\text{Ca}$
c.	8	9	oxygen-17	$^{17}_8\text{O}$
d.	26	31	iron-57	$^{57}_{26}\text{Fe}$
e.	30	34	zinc-64	$^{64}_{30}\text{Zn}$
f.	80	124	mercury-204	$^{204}_{80}\text{Hg}$

17. **Challenge** An atom has a mass number of 55. Its number of neutrons is the sum of its atomic number and 5. How many protons, neutrons, and electrons does this atom have? What is the identity of this atom?

25 protons, 25 electrons, 30 neutrons, manganese

18. Boron (B) has two naturally occurring isotopes: boron-10 (abundance = 19.8%, mass = 10.013 amu), boron-11 (abundance = 80.2%, mass = 11.009 amu). Calculate the atomic mass of boron.

$$\text{B-10 } 10.013 \text{ amu} \times 0.198 = 1.98 \text{ amu}$$

$$\text{B-11 } 11.009 \text{ amu} \times 0.802 = 8.83 \text{ amu}$$

$$\text{Atomic mass } 1.98 \text{ amu} + 8.83 \text{ amu} = 10.81 \text{ amu}$$

19. **Challenge** Nitrogen has two naturally occurring isotopes, N-14 and N-15. Its atomic mass is 14.007. Which isotope is more abundant? Explain your answer.

N-14 is more abundant because the atomic mass is closer to 14 than 15.

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mar 27 2019 gjenerali i juntes ushtarake ne shqiperi lleshi siguron francezet se do ti shtrengoje shqipetaret te mos kerkojne azil ne france përgjigju sondazhi i ditës kërcënimet nga krimi a duhet forcuar mbrojtja e prokurorëve e gjyqtarëve po jo nuk e di lajme të tjera operacioni tempulli ish ministri i brendshëm në kontrast si

azili në francë kërkohet shpallja e shqipëris syri - May 07 2022

16 04 14 02 2020 a a a shqipëria dhe dy vende të afrikës perëndimore janë propozuar për t u hequr nga lista e vendeve të sigurta të origjinës nga ojf ja më e madhe për mbrojtjen e të drejtave të refugjatëve në francë

[prijava za azil u francuskoj women for women france](#) - Feb 16 2023

may 22 2023 prijava za azil je postupak koji omogućava ljudima rođenim van francuske da dobiju status izbeglice ili supsidijarnu zaštitu od progona ili rizika od progona u svojoj zemlji porekla potvrđeno od strane alexandra lachowsky na dan 22 05 2023 prava boravka u francuskoj prijava za azil u francuskoj

aplikimet për azil në francë janë rritur për 7 3 për qind - Apr 18 2023

21 janar 2020 23 55 aplikimet për azil në francë u rritën me 7 3 në vitin 2019 krahasuar me një vit më parë sipas të dhënave të publikuara nga zyra ranceze për mbrojtjen e refugjatëve ofpra numri i aplikantëve për azil arriti në 132 mijë e 614 persona