

## Chapter 6 – Study Guide

Test Date: Friday, December 19, 2014

Use the following to study:

Textbook – Ch. 6

Ch. 6 homework packet

### Chemical Bonding:

- **Four types of bonds:**
  - *Nonpolar Covalent Bond* ( $\Delta EN$  (difference in electronegativity) = < 0.3)
    - EQUAL SHARING of electrons between the 2 bonding nonmetal atoms
  - *Polar Covalent Bond* ( $\Delta EN = 0.4-1.7$ )
    - UNEQUAL SHARING of electrons between the 2 bonding nonmetal atoms (because EN is much higher in one atom, the electrons are pulled closer to the atom in the bond with the higher EN value – like an uneven tug of war)
  - *Ionic Bond* ( $\Delta EN = > 1.7$ )
    - Difference in the EN values for the 2 bonding atoms (metal & nonmetal) is so large that the atom with the lower EN value ends up completely GIVING UP/TRANSFERRING ELECTRONS to the atom with the greater EN value
    - Ionic bond results due to the attraction between the POSITIVE cation (formed by losing electrons) and the NEGATIVE anion (formed by gaining electrons)
  - *Metallic Bond*
    - Bonding between two metallic atoms which occurs due to the attraction between a “sea of electrons” and the nuclei of bonded metal atoms
- \* Must memorize/know the difference in electronegativity ranges for each type of bond
  - \*\* Note: you should still know the definition of electronegativity (EN)
- **Properties of the Four Types of Chemical Bonds**
  - See the table at the bottom of Packet pg.5
  - You MUST be able to compare all four examples – NaCl (ionic), Sucrose (Polar), Stearic Acid/Wax (nonpolar), and Aluminum & Mercury (metallic) – in terms of melting point, conductivity of solid, solubility in water, conductivity in water (when applicable), solubility in hexane, hardness (ease with which it is scratched), and brittleness
- **Why & How do chemical bonds form?**
  - Why → because all atoms want to be at their LOWEST POTENTIAL ENERGY; for nearly all atoms (except the noble gases) this happens only when an atom chemically bonds to other atoms!!!
  - How → ATTRACTIVE and REPULSIVE forces
    - *Attractive forces* – positively charged nucleus of one atom is attracted to electrons in another atom nearby (because they have opposite charges)
    - *Repulsive forces* – 2 nuclei or 2 sets of electron clouds from nearby atoms repel one another (because they have the same charge)
  - **Bond Length**
    - ideal distance between 2 atoms that gives the lowest potential energy of the atoms by balancing out attractive & repulsive forces)
  - **Bond Energy**
    - Energy required to break apart a chemical bond (the higher the bond energy, the stronger the bond)

# Chapter 6 Study Guide Chemistry

**Didier Musso**



## Chapter 6 Study Guide Chemistry:

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**Student Study Guide to Accompany Petrucci's General Chemistry, 3rd. Ed** Robert K. Wismer, 1982

**Study Guide [to Accompany] General Chemistry** Dixie Goss, 2002

*Study Guide for General Chemistry and College Chemistry, Eighth Editions by Holtzclaw and Robinson* Norman E. Griswold, Henry Fuller Holtzclaw, 1988

**Study Guide for Chemistry, Third Edition [by] Steven S. Zumdahl** Paul B. Kelter, 1993

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*Fundamentals of Organic Chemistry, Textbook, Study Guide and Solutions Manual* T. W. Graham Solomons,1996-08-15 A realistic approach to the study of mechanisms The book addresses real functional group chemistry with an emphasis on the biological environmental and medical applications of organic chemistry

*Study Guide and Solutions Manual to Accompany Fundamentals of Organic Chemistry* McMurry,Susan McMurry,1990

**Robinson Chemistry Study Guide** Robinson,1992

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## **Table of Contents Chapter 6 Study Guide Chemistry**

1. Understanding the eBook Chapter 6 Study Guide Chemistry
  - The Rise of Digital Reading Chapter 6 Study Guide Chemistry
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 6 Study Guide Chemistry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 6 Study Guide Chemistry
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 6 Study Guide Chemistry
  - Personalized Recommendations
  - Chapter 6 Study Guide Chemistry User Reviews and Ratings
  - Chapter 6 Study Guide Chemistry and Bestseller Lists
5. Accessing Chapter 6 Study Guide Chemistry Free and Paid eBooks
  - Chapter 6 Study Guide Chemistry Public Domain eBooks
  - Chapter 6 Study Guide Chemistry eBook Subscription Services

- Chapter 6 Study Guide Chemistry Budget-Friendly Options
- 6. Navigating Chapter 6 Study Guide Chemistry eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 6 Study Guide Chemistry Compatibility with Devices
  - Chapter 6 Study Guide Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 6 Study Guide Chemistry
  - Highlighting and Note-Taking Chapter 6 Study Guide Chemistry
  - Interactive Elements Chapter 6 Study Guide Chemistry
- 8. Staying Engaged with Chapter 6 Study Guide Chemistry
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 6 Study Guide Chemistry
- 9. Balancing eBooks and Physical Books Chapter 6 Study Guide Chemistry
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 6 Study Guide Chemistry
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 6 Study Guide Chemistry
  - Setting Reading Goals Chapter 6 Study Guide Chemistry
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 6 Study Guide Chemistry
  - Fact-Checking eBook Content of Chapter 6 Study Guide Chemistry
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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