



COVALENT BONDING

Practice Problems

Do your best to solve each of the following problems.

SECTION 8.1 MOLECULAR COMPOUNDS

- Classify each of the following as either an ionic or a molecular compound.
 - NaCl
 - H_2O
 - CaF_2
 - C_2H_6
 - KNO_3
 - H_2S
 - MgO
 - C_2H_4

- Classify each of the following as either an ionic or a molecular compound.
 - C_2H_6
 - H_2O
 - CaF_2
 - KNO_3
 - MgO
 - C_2H_4
 - NaCl
 - H_2S
- Write the formula for each of the following molecular compounds.
 - Carbon dioxide
 - Carbon monoxide
 - Diphosphorus pentoxide
 - Diphosphorus trioxide
 - Diphosphorus pentasulfide
 - Diphosphorus trisulfide
 - Diphosphorus tetrasulfide
 - Diphosphorus hexasulfide
 - Diphosphorus heptasulfide
 - Diphosphorus octasulfide
 - Diphosphorus nonasulfide
 - Diphosphorus decasulfide
 - Diphosphorus undecasulfide
 - Diphosphorus dodecasulfide
 - Diphosphorus trisulfide
 - Diphosphorus tetrasulfide
 - Diphosphorus pentasulfide
 - Diphosphorus hexasulfide
 - Diphosphorus heptasulfide
 - Diphosphorus octasulfide
 - Diphosphorus nonasulfide
 - Diphosphorus decasulfide
 - Diphosphorus undecasulfide
 - Diphosphorus dodecasulfide

SECTION 8.2 THE NATURE OF COVALENT BONDING

- Draw the Lewis structure for hydrogen fluoride. $\text{H}-\text{F}$
- Draw the Lewis structure for carbon dioxide. $\text{O}=\text{C}=\text{O}$
- Draw the Lewis structure for methane. C with four H atoms bonded to it.
- Draw the Lewis structure for acetylene. $\text{H}-\text{C}\equiv\text{C}-\text{H}$
- Draw the Lewis structure for hydrogen cyanide. $\text{H}-\text{C}\equiv\text{N}$



SECTION 8.3 BONDING IN THE LHS

- Write the Lewis structure for each of the following molecular compounds.
 - C_2H_6
 - H_2O
 - CaF_2
 - KNO_3
 - MgO
 - C_2H_4
 - NaCl
 - H_2S

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