



# COVALENT BONDING

## Practice Problems

Do your best to solve each of the following problems.

### SECTION 8.1 MOLECULAR COMPOUNDS

- Classify each of the following compounds as ionic or covalent.
  - $\text{Na}_2\text{SO}_4$
  - $\text{C}_2\text{H}_6$
  - $\text{H}_2\text{O}$
  - $\text{KNO}_3$
  - $\text{CaCl}_2$
  - $\text{H}_2\text{S}$
  - $\text{H}_2\text{SO}_4$
  - $\text{H}_2\text{SO}_3$
  - $\text{H}_2\text{SO}_2$
  - $\text{H}_2\text{SO}$
  - $\text{H}_2\text{S}_2$
  - $\text{H}_2\text{S}_3$
  - $\text{H}_2\text{S}_4$
  - $\text{H}_2\text{S}_5$
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  - $\text{H}_2\text{S}_{50}$

### SECTION 8.2 THE NATURE OF COVALENT BONDING

- Show covalent bonding in hydrogen fluoride.  $\text{H}-\text{F}$
- Show covalent bonding in methane.  $\text{C}$  with four  $\text{H}$  atoms around it.
- Show covalent bonding in carbon dioxide.  $\text{C}$  with two  $\text{O}$  atoms around it.
- Show covalent bonding in ethane.  $\text{H}-\text{C}-\text{C}-\text{H}$
- Show covalent bonding in acetylene.  $\text{H}-\text{C}\equiv\text{C}-\text{H}$
- Draw Lewis structures for the following molecules:
  - $\text{H}_2\text{O}$
  - $\text{H}_2\text{S}$
  - $\text{H}_2\text{Se}$
  - $\text{H}_2\text{Te}$
  - $\text{H}_2\text{O}_2$
  - $\text{H}_2\text{S}_2$
  - $\text{H}_2\text{S}_3$
  - $\text{H}_2\text{S}_4$
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  - $\text{H}_2\text{S}_{50}$

### SECTION 8.3 BONDING: THE ORBITALS

- Which orbital and sub-orbital are represented by  $\text{p}_x$ ,  $\text{p}_y$ ,  $\text{p}_z$ ,  $\text{d}_{xy}$ ,  $\text{d}_{yz}$ ,  $\text{d}_{zx}$ ,  $\text{d}_{x^2-y^2}$ , and  $\text{d}_{z^2}$ ?
- Which orbital and sub-orbital are represented by  $\text{p}_x$ ,  $\text{p}_y$ ,  $\text{p}_z$ ,  $\text{d}_{xy}$ ,  $\text{d}_{yz}$ ,  $\text{d}_{zx}$ ,  $\text{d}_{x^2-y^2}$ , and  $\text{d}_{z^2}$ ?

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**R Sandford**



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