

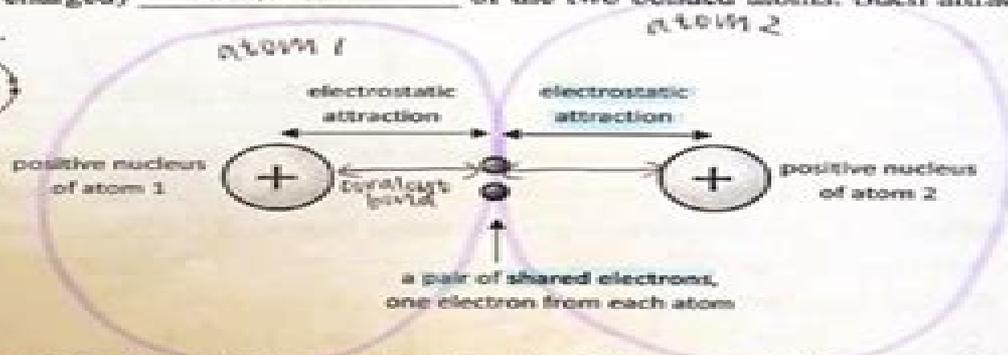
Chapter 8: Covalent Bonding

Students should:

- 1.14 know what is meant by the term *molecule*
- 1.44 know that a covalent bond is formed between atoms by the sharing of a pair of electrons
- 1.45 understand covalent bonds in terms of electrostatic attractions
- 1.46 understand how to use dot-and-cross diagrams to represent covalent bonds in:
 - diatomic molecules, including hydrogen, oxygen, nitrogen, halogens and hydrogen halides
 - inorganic molecules including water, ammonia and carbon dioxide
 - organic molecules containing up to two carbon atoms, including methane, ethane, ethene and those containing halogen atoms
- 1.47 explain why substances with simple molecular structures are gases or liquids, or solids with low melting and boiling points
the term intermolecular forces of attraction can be used to represent all forces between molecules
- 1.48 explain why the melting and boiling points of substances with simple molecular structures increase, in general, with increasing relative molecular mass
- 1.49 explain why substances with giant covalent structures are solids with high melting and boiling points
- 1.50 explain how the structures of diamond, graphite and C₆₀ fullerene influence their physical properties, including electrical conductivity and hardness
- 1.51 know that covalent compounds do not usually conduct electricity

8.1 Covalent bonding

A covalent bond is resulted when non-metal atoms combine with non-metal atoms by sharing of outermost shell electrons. Strong, directional electrostatic attraction is formed between the (negatively charged) shared electrons and the (positively charged) nuclei of the two bonded atoms. Such attraction is called a



The bonded atoms can then attain an octet or duplet of electrons and become stable (i.e. each atom attains the stable electronic structure of its nearest noble gas).

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Table of Contents Chapter 8 Covalent Bonding Work Answers Prentice Hall

1. Understanding the eBook Chapter 8 Covalent Bonding Work Answers Prentice Hall
 - The Rise of Digital Reading Chapter 8 Covalent Bonding Work Answers Prentice Hall
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 8 Covalent Bonding Work Answers Prentice Hall
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 8 Covalent Bonding Work Answers Prentice Hall
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 8 Covalent Bonding Work Answers Prentice Hall
 - Personalized Recommendations
 - Chapter 8 Covalent Bonding Work Answers Prentice Hall User Reviews and Ratings

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- Chapter 8 Covalent Bonding Work Answers Prentice Hall and Bestseller Lists
- 5. Accessing Chapter 8 Covalent Bonding Work Answers Prentice Hall Free and Paid eBooks
 - Chapter 8 Covalent Bonding Work Answers Prentice Hall Public Domain eBooks
 - Chapter 8 Covalent Bonding Work Answers Prentice Hall eBook Subscription Services
 - Chapter 8 Covalent Bonding Work Answers Prentice Hall Budget-Friendly Options
- 6. Navigating Chapter 8 Covalent Bonding Work Answers Prentice Hall eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 8 Covalent Bonding Work Answers Prentice Hall Compatibility with Devices
 - Chapter 8 Covalent Bonding Work Answers Prentice Hall Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 8 Covalent Bonding Work Answers Prentice Hall
 - Highlighting and Note-Taking Chapter 8 Covalent Bonding Work Answers Prentice Hall
 - Interactive Elements Chapter 8 Covalent Bonding Work Answers Prentice Hall
- 8. Staying Engaged with Chapter 8 Covalent Bonding Work Answers Prentice Hall
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 8 Covalent Bonding Work Answers Prentice Hall
- 9. Balancing eBooks and Physical Books Chapter 8 Covalent Bonding Work Answers Prentice Hall
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 8 Covalent Bonding Work Answers Prentice Hall
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 8 Covalent Bonding Work Answers Prentice Hall
 - Setting Reading Goals Chapter 8 Covalent Bonding Work Answers Prentice Hall
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 8 Covalent Bonding Work Answers Prentice Hall
 - Fact-Checking eBook Content of Chapter 8 Covalent Bonding Work Answers Prentice Hall
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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