

## Chapter Test B

### Chapter: Chemical Equations and Reactions

Match a name space (possibly writing the name of the react or product molecule) completed with reactants/products atoms/molecules.

- The product(s) of a chemical reaction will be represented by a double arrow pointing towards the right. The substance(s) on the left side of the reaction is/are
  - reactant(s)**
  - product(s)
  - solvent
- An ionic chemical equation is represented by adding the
  - charges**
  - charges of the particles
  - number of particles
- An ionic equation that is written as  $aq$  indicates the substance is
  - solid
  - dissolved in water**
  - liquid
  - volatile
- The activity for precipitation reactions are used to predict if the
  - solubility of the reactants and products for the reaction will be the same
  - precipitate will form
  - no precipitate will form
  - solubility of the reactants is different from the solubility of the products**
- The coefficients for chemical equations
  - indicate the number of moles of each substance that react**
  - show the number of grams of each substance that react
  - show the number of molecules of the substances
  - show the relative proportions of each atom
- If you're to balance the chemical  $Fe + O_2 \rightarrow Fe_2O_3$  with one  $Fe$  atom, then
  - 2** will replace  $O_2$  in the equation
  - 3** will replace  $O_2$  in the equation
  - 1 will replace  $O_2$  in the equation
  - 2 will replace  $O_2$  in the equation

# Chapter 8 Test B

**Shasha Hu**



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