

## Section 12.1

### OBJECTIVE 1

Write complete balanced redox reactions that are indicated.

1.  $\text{Fe}^{2+} + \text{H}_2\text{O}_2 \rightarrow \text{Fe}^{3+} + \text{H}_2\text{O}$  (acidic solution)
  - (a) identify the oxidizing and reducing agents
  - (b) balance the reaction in acidic solution
  - (c) identify the oxidized and reduced species
  - (d) identify the oxidizing and reducing agents
2.  $\text{Fe}^{2+} + \text{H}_2\text{O}_2 \rightarrow \text{Fe}^{3+} + \text{H}_2\text{O}$  (basic solution)
  - (a) identify the oxidizing and reducing agents
  - (b) balance the reaction in basic solution
  - (c) identify the oxidized and reduced species
  - (d) identify the oxidizing and reducing agents
3.  $\text{Fe}^{2+} + \text{H}_2\text{O}_2 \rightarrow \text{Fe}^{3+} + \text{H}_2\text{O}$  (neutral solution)
  - (a) identify the oxidizing and reducing agents
  - (b) balance the reaction in neutral solution
  - (c) identify the oxidized and reduced species
  - (d) identify the oxidizing and reducing agents
4.  $\text{Fe}^{2+} + \text{H}_2\text{O}_2 \rightarrow \text{Fe}^{3+} + \text{H}_2\text{O}$  (acidic solution)
  - (a) identify the oxidizing and reducing agents
  - (b) balance the reaction in acidic solution
  - (c) identify the oxidized and reduced species
  - (d) identify the oxidizing and reducing agents

5. Write the net ionic equation for the following redox reaction:



a. Balance the equation:



b. Write the cell notation for the reaction.

c. Calculate the cell potential for the reaction.

d. Write the half-cell reactions for the reaction.

e. Calculate the Gibbs free energy for the reaction.

# Chapter 9 Review Stoichiometry Section 1 Answers

**Ralph Thomas Myers**



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