

Section 12.1

OBJECTIVE 1

Write complete balanced redox reactions that are indicated.

1. $\text{Fe}^{2+} + \text{MnO}_4^- + \text{H}^+$
 - (a) Write the balanced redox equation assuming that reaction is given in acidic solution.
 - (b) Calculate standard cell potential of reaction and ΔG° .
 - (c) Calculate $\log K$ for reaction in acidic solution.
 - (d) Calculate $\log K$ for reaction in basic solution.
2. $\text{Fe}^{2+} + \text{H}_2\text{O}_2 + \text{H}^+$
 - (a) Write the balanced redox equation assuming that reaction is given in acidic solution.
 - (b) Calculate standard cell potential of reaction and ΔG° .
 - (c) Calculate $\log K$ for reaction in acidic solution.
 - (d) Calculate $\log K$ for reaction in basic solution.
3. $\text{Fe}^{2+} + \text{H}_2\text{O}_2 + \text{OH}^-$
 - (a) Write the balanced redox equation assuming that reaction is given in basic solution.
 - (b) Calculate standard cell potential of reaction and ΔG° .
 - (c) Calculate $\log K$ for reaction in basic solution.
 - (d) Calculate $\log K$ for reaction in acidic solution.
4. $\text{Fe}^{2+} + \text{H}_2\text{O}_2 + \text{H}^+$
 - (a) Write the balanced redox equation assuming that reaction is given in acidic solution.
 - (b) Calculate standard cell potential of reaction and ΔG° .
 - (c) Calculate $\log K$ for reaction in acidic solution.
 - (d) Calculate $\log K$ for reaction in basic solution.

5. Write the net redox reaction for the following half-cell reactions:



(a) Balance the equation:



(b) Write the cell notation for this cell.

(c) Calculate $\log K$ for this reaction assuming standard conditions.

(d) Calculate standard potential of MnO_2 from standard potential of Mn^{2+} .

(e) Calculate standard potential of MnO_2 from standard potential of H_2O .

Chapter 9 Section 1 Review Stoichiometry Answers

Raffaella Di Napoli



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