

## SOLVED QUESTION BANK

## Objective / Very Short Answer Type Questions [1 Mark]

1. The half life period of first order reaction is 1386 seconds. The specific rate constant of the reaction is

- (a)  $0.5 \times 10^{-2} \text{ s}^{-1}$  (b)  $0.5 \times 10^{-3} \text{ s}^{-1}$   
(c)  $5.0 \times 10^{-2} \text{ s}^{-1}$  (d)  $5.0 \times 10^{-3} \text{ s}^{-1}$

Ans. (b)  $t_{1/2} = \frac{0.693}{k}$   
 $\Rightarrow k = \frac{0.693}{1386} = 0.5 \times 10^{-3} \text{ s}^{-1}$

2. The rate constant of a reaction  $A \longrightarrow B$  is  $0.6 \times 10^{-3}$  mole per second. If the concentration of [A] is 5 M, then what will be concentration of [B] after 20 minutes?

- (a) 0.36 M (b) 0.72 M  
(c) 1.08 M (d) 3.60 M

Ans. (b)  $k = \frac{[R]_0 - [R]}{t}$  for zero order.  
 $\Rightarrow 0.6 \times 10^{-3} \times 20 \times 60 = [R]_0 - [R]$   
 $\Rightarrow 0.72 \text{ M}$

Amount of [B] produced =  $[R]_0 - [R] = 0.72 \text{ M}$

3. Which of the following is affected by catalyst?

- (a)  $\Delta H$  (b)  $\Delta S$   
(c)  $\Delta G$  (d)  $E_a$

Ans. (d) Catalyst affects activation energy. It does not affect,  $\Delta H$ ,  $\Delta S$ ,  $\Delta G$ .

4. A first order reaction has specific reaction rate  $10^{-2} \text{ s}^{-1}$ . How much time it will take for 20g of reactant to reduce to 5g?

- (a) 138.6 s (b) 346.5 s  
(c) 693.0 s (d) 238.6 s

Ans. (a)  $t_{3/4} = \frac{1.386}{k} = \frac{1.386}{10^{-2}} = 138.6 \text{ s}$

5. The activation energy of a reaction can be determined from the slope of which of the following graph:

- (a)  $\ln k$  Vs  $\frac{1}{T}$  (b)  $\frac{T}{\ln k}$  Vs  $\frac{1}{T}$   
(c)  $\ln k$  Vs  $T$  (d)  $\frac{\ln k}{T}$  Vs  $T$

Ans. (a) When we plot  $\ln k$  Vs  $\frac{1}{T}$ , slope of line =  $-\frac{E_a}{R}$ .

6. Mechanism of a hypothetical reaction  $X_2 + Y_2 \longrightarrow 2XY$  is given below:

- (a)  $X_2 \rightleftharpoons X + X$  (fast)  
(b)  $X + Y_2 \longrightarrow XY + Y$  (slow)  
(c)  $X + Y \longrightarrow XY$  (fast)

The overall order of reaction is

- (a) 2 (b) 0 (c) 1.5 (d) 1

Ans. (c) rate =  $k[X][Y_2]$ ,  $K = \frac{[X]^2}{[X_2]} \Rightarrow [X] = \sqrt{K[X_2]}$   
 rate =  $k \times K^{1/2} [X_2]^{1/2} [Y_2] = k^1 [X_2]^{1/2} [Y_2]^1$   
 Order =  $\frac{1}{2} + 1 = \frac{3}{2} = 1.5$

7. The rate of first order reaction is  $0.04 \text{ mol L}^{-1} \text{ s}^{-1}$  at 10 sec, and  $0.03 \text{ mol L}^{-1}$  at 20 seconds after initiation of the reaction.  $t_{1/2}$  of reaction is

- (a) 44.1 s (b) 54.1 s (c) 24.1 s (d) 34.1 s

Ans. (c)  $k = \frac{2.303}{20 - 10} \log \frac{0.04}{0.03} = \frac{2.303}{10} \log \frac{4}{3}$   
 $= \frac{2.303}{10} (\log 4 - \log 3)$   
 $= \frac{2.303}{10} (0.6021 - 0.4771) = \frac{2.303}{10} \times 0.1249$   
 $t_{1/2} = \frac{0.693}{k} = \frac{0.693 \times 10}{2.303 \times 0.1249} = \frac{3.010}{0.1249} = 24.1 \text{ s}$

8. If the initial concentration of reactant is doubled  $t_{1/2}$  is also doubled, the order of reaction is

- (a) zero (b) 1 (c) 2 (d) 3

Ans. (a) In zero order of reaction  $t_{1/2} \propto [R]_0$

9. For a reaction taking place in three steps, the overall rate constant  $k = \frac{k_1 k_2}{k_3}$ . If  $E_{a1}$ ,  $E_{a2}$  and  $E_{a3}$  are 40, 50 and 60 kJ  $\text{mol}^{-1}$ , the overall  $E_a$  is

- (a) 30 (b) 40 (c) 60 (d) 50

Ans. (a)  $k = A e^{-E_a/RT}$   
 $k = \frac{k_1 k_2}{k_3}$   
 $\Rightarrow A e^{-E_a/RT} = \frac{A e^{-E_{a1}/RT} \times A e^{-E_{a2}/RT}}{A e^{-E_{a3}/RT}}$   
 $e^{-E_a/RT} = \frac{e^{-E_{a1}/RT} \times e^{-E_{a2}/RT}}{e^{-E_{a3}/RT}}$

Taking  $\ln$  on both side we get

$$\frac{E_a}{RT} = -E_{a1} - E_{a2} + E_{a3}$$

$$E_a = E_{a1} + E_{a2} - E_{a3}$$

$$E_a = 40 + 50 - 60 = 30 \text{ kJ mol}^{-1}$$

10. If conc. of reactant 'A' is increased 10 times and rate of reaction becomes 100 times. What is order with respect to 'A'?

- (a) 1 (b) 2 (c) 3 (d) 4

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