

Experiment 18-Chemical Reactions Lab Report

Name: _____

Section: _____

For this lab, you need to balance each equation, and classify the reaction according to the types of reactions listed in the lab handout using C (combination), D (decomposition), Cu (combustion), S (single replacement) and DR (double replacement). Include what you observed below each reaction.

- | | <u>Type of Reaction</u> |
|--|-------------------------|
| 1. $\underline{\hspace{1cm}} \text{CaCl}_2 (\text{aq}) + \underline{\hspace{1cm}} \text{Na}_3\text{PO}_4 (\text{aq}) \rightarrow \underline{\hspace{1cm}} \text{Ca}_3(\text{PO}_4)_2 (\text{s}) + \underline{\hspace{1cm}} \text{NaCl} (\text{aq})$ | _____ |
| Observations: | |
| 2. $\underline{\hspace{1cm}} \text{CuSO}_4 (\text{H}_2\text{O})_5 + \text{heat} \rightarrow \underline{\hspace{1cm}} \text{CuSO}_4 (\text{s}) + \underline{\hspace{1cm}} \text{H}_2\text{O} (\text{g})$ | _____ |
| Observations: | |
| $\underline{\hspace{1cm}} \text{CuSO}_4 (\text{s}) + \underline{\hspace{1cm}} \text{H}_2\text{O} (\text{l}) \rightarrow \underline{\hspace{1cm}} \text{CuSO}_4 (\text{H}_2\text{O})_5 (\text{s})$ | _____ |
| Observations: | |
| 3. $\underline{\hspace{1cm}} \text{Zn} (\text{s}) + \underline{\hspace{1cm}} \text{HCl} (\text{aq}) \rightarrow \underline{\hspace{1cm}} \text{ZnCl}_2 (\text{aq}) + \underline{\hspace{1cm}} \text{H}_2 (\text{g})$ | _____ |
| Observations: | |
| $\underline{\hspace{1cm}} \text{H}_2 (\text{g}) + \underline{\hspace{1cm}} \text{O}_2 (\text{g}) \rightarrow \underline{\hspace{1cm}} \text{H}_2\text{O} (\text{g})$ The "popping" reaction induced with the match. | _____ |
| Observations: | |
| 4. $\underline{\hspace{1cm}} \text{HCl} (\text{aq}) + \underline{\hspace{1cm}} \text{Na}_2\text{CO}_3 (\text{aq}) \rightarrow \underline{\hspace{1cm}} \text{CO}_2 (\text{g}) + \underline{\hspace{1cm}} \text{NaCl} (\text{aq}) + \underline{\hspace{1cm}} \text{H}_2\text{O} (\text{l})$ | _____ |
| Observations: | |
| 5. $\underline{\hspace{1cm}} \text{H}_2\text{O}_2 (\text{aq}) \rightarrow \underline{\hspace{1cm}} \text{O}_2 (\text{g}) + \underline{\hspace{1cm}} \text{H}_2\text{O} (\text{l})$ (the MnO_2 is a catalyst, it does not react) | _____ |
| Observations: | |
| 6. $\underline{\hspace{1cm}} \text{HCl} (\text{aq}) + \underline{\hspace{1cm}} \text{NaOH} (\text{aq}) \rightarrow \underline{\hspace{1cm}} \text{NaCl} (\text{aq}) + \underline{\hspace{1cm}} \text{H}_2\text{O} (\text{l})$ | _____ |
| Observations: | |
| 7. $\underline{\hspace{1cm}} \text{CaCl}_2 (\text{aq}) + \underline{\hspace{1cm}} \text{Na}_2\text{CO}_3 (\text{aq}) \rightarrow \underline{\hspace{1cm}} \text{NaCl} (\text{aq}) + \underline{\hspace{1cm}} \text{CaCO}_3 (\text{s})$ | _____ |
| Observations: | |
| 8. $\underline{\hspace{1cm}} \text{Zn} (\text{s}) + \underline{\hspace{1cm}} \text{CuSO}_4 (\text{aq}) \rightarrow \underline{\hspace{1cm}} \text{ZnSO}_4 (\text{aq}) + \underline{\hspace{1cm}} \text{Cu} (\text{s})$ | _____ |
| Observations: | |
| 9. $\underline{\hspace{1cm}} \text{Cu} (\text{s}) + \underline{\hspace{1cm}} \text{ZnSO}_4 (\text{aq}) \rightarrow \underline{\hspace{1cm}} \text{CuSO}_4 (\text{aq}) + \underline{\hspace{1cm}} \text{Zn} (\text{s})$ | _____ |
| Observations: | |
| 10. $\underline{\hspace{1cm}} \text{CaCl}_2 (\text{aq}) + \underline{\hspace{1cm}} \text{NaNO}_3 (\text{aq}) \rightarrow \underline{\hspace{1cm}} \text{Ca}(\text{NO}_3)_2 (\text{aq}) + \underline{\hspace{1cm}} \text{NaCl} (\text{aq})$ | _____ |
| Observations: | |

Chemical Reactions Lab Answers

Thomas Griffiths



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