

Chapter Test II

A. Matching

Match each item in Column I with the correct description in Column II. Write the letter of the correct description.

Column I

1. An oxidation-reduction reaction in which an element is oxidized and another element is reduced.
2. An equation of a redox reaction that is balanced for both mass and charge.
3. The reduction of a quantity to a lower oxidation state.
4. The ratio of the actual yield to the theoretical yield, expressed as a percent.
5. An equation that is balanced for mass, but not for charge.
6. An equation that is balanced for mass and charge.

Column II

- a. percent yield
- b. limiting reagent
- c. oxidation state
- d. half-reaction
- e. balanced equation
- f. redox reaction

B. Multiple Choice

Choose the best answer for each question.

1. Which of the following is not a redox reaction?
 - a. $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 - b. $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$
 - c. $\text{H}_2 + \text{Cu}^{2+} \rightarrow \text{Cu} + 2\text{H}^+$
 - d. $\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
2. Which of the following is not a redox reaction?
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 - c. $\text{H}_2 + \text{Cu}^{2+} \rightarrow \text{Cu} + 2\text{H}^+$
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 - d. $\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$

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Léo Pomar



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