

## Chemical Bonding

### SECTION 3

SHORT ANSWER Answer the following questions in the space provided.

- A** The notation for sodium chloride, NaCl, stands for one  
(a) formula unit. (c) crystal.  
(b) molecule. (d) atom.
- D** In a crystal of an ionic compound, each cation is surrounded by a number of  
(a) molecules. (c) dipoles.  
(b) positive ions. (d) negative ions.
- B** Compared with the neutral atoms involved in the formation of an ionic compound, the crystal lattice that results is  
(a) higher in potential energy. (c) equal in potential energy.  
(b) lower in potential energy. (d) unstable.
- B** The lattice energy of compound A is greater in magnitude than that of compound B. What can be concluded from this fact?  
(a) Compound A is not an ionic compound.  
(b) It will be more difficult to break the bonds in compound A than those in compound B.  
(c) Compound B has larger crystals than compound A.  
(d) Compound A has larger crystals than compound B.
- B** The forces of attraction between molecules in a molecular compound are generally  
(a) stronger than the attractive forces among formula units in ionic bonding.  
(b) weaker than the attractive forces among formula units in ionic bonding.  
(c) approximately equal to the attractive forces among formula units in ionic bonding.  
(d) equal to zero.
- Describe the force that holds two ions together in an ionic bond.

**Very strong overall attraction between positive and negative charges**

- What type of energy best represents the strength of an ionic bond?

**Lattice energy**

- What types of bonds are present in an ionic compound that contains a polyatomic ion?

# Chemistry Chapter 6 Review Answers

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## **Table of Contents Chemistry Chapter 6 Review Answers**

1. Understanding the eBook Chemistry Chapter 6 Review Answers
  - The Rise of Digital Reading Chemistry Chapter 6 Review Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 6 Review Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Chapter 6 Review Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 6 Review Answers
  - Personalized Recommendations
  - Chemistry Chapter 6 Review Answers User Reviews and Ratings
  - Chemistry Chapter 6 Review Answers and Bestseller Lists
5. Accessing Chemistry Chapter 6 Review Answers Free and Paid eBooks
  - Chemistry Chapter 6 Review Answers Public Domain eBooks

- Chemistry Chapter 6 Review Answers eBook Subscription Services
- Chemistry Chapter 6 Review Answers Budget-Friendly Options
- 6. Navigating Chemistry Chapter 6 Review Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Chapter 6 Review Answers Compatibility with Devices
  - Chemistry Chapter 6 Review Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Chapter 6 Review Answers
  - Highlighting and Note-Taking Chemistry Chapter 6 Review Answers
  - Interactive Elements Chemistry Chapter 6 Review Answers
- 8. Staying Engaged with Chemistry Chapter 6 Review Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Chapter 6 Review Answers
- 9. Balancing eBooks and Physical Books Chemistry Chapter 6 Review Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Chapter 6 Review Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Chapter 6 Review Answers
  - Setting Reading Goals Chemistry Chapter 6 Review Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Chapter 6 Review Answers
  - Fact-Checking eBook Content of Chemistry Chapter 6 Review Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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