

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

## CHAPTER 13 STUDY GUIDE FOR CONTENT MASTERY

## States of Matter

## Section 13.1 Gases

In your textbook, read about the kinetic molecular theory.

Complete each statement.

- The kinetic molecular theory describes the behavior of gases in terms of particles in motion.
- The kinetic molecular theory makes the following assumptions:
  - In a sample of a gas, the volume of the gas particles themselves is very small compared to the volume of the sample.
  - Between gas particles and the walls, there are no significant attractive or repulsive forces between gas particles.
  - Gas particles are in constant and random motion.
  - The collisions between gas particles are elastic, that is, no kinetic energy is lost.
- The kinetic energy of a particle is represented by the equation  $KE = \frac{1}{2}mv^2$ .
- Temperature is a measure of the average kinetic energy of the particles in a sample of matter.

In your textbook, read about explaining the behavior of gases.

For each statement below, write true or false.

- |              |  |
|--------------|--|
| <u>True</u>  | 1. Gases are less dense than solids because there is a lot of space between the particles of a gas.                              |
| <u>True</u>  | 2. The random motion of gas particles causes a gas to expand until it fills its container.                                       |
| <u>False</u> | 3. The density of a gas decreases as it is compressed.   |
| <u>True</u>  | 4. A gas can flow into a space occupied by another gas.  |
| <u>True</u>  | 5. The diffusion of a gas is caused by the random motion of the particles of the gas.  |
| <u>False</u> | 6. Lighter gas particles diffuse less rapidly than do heavier gas particles.   |
| <u>True</u>  | 7. During effusion, a gas escapes through a tiny opening into a vacuum.  |
| <u>False</u> | 8. Graham's law of effusion states that the rate of effusion for a gas is directly related to the square root of its molar mass. |

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Chemistry: Matter and Change • Chapter 13 79

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

## CHAPTER 13 STUDY GUIDE FOR CONTENT MASTERY

## Section 13.1 continued

In your textbook, read about gas pressure.

Circle the letter of the choice that best completes the statement or answers the question.

- Pressure is defined as force per unit
  - area
  - time
  - mass
  - volume
- What is an instrument designed to measure atmospheric pressure?
  - barometer
  - manometer
  - physoconimeter
  - densimeter
- The height of the liquid in a barometer is affected by all of the following EXCEPT the
  - altitude
  - density of the liquid in the column
  - atmospheric pressure
  - diameter of the column tube.
- The pressure of the gas in a manometer is directly related to which of the following
  - height of the mercury column in the closed-end arm
  - height of the mercury column in the open-end arm
  - $h + 76$
  - $h - 76$
- One atmosphere is equal to a pressure of
  - 76 mm Hg
  - 101.3 kPa
  - 147 psi
  - 760 mm
- The partial pressure of a gas depends on all of the following EXCEPT the
  - concentration of the gas
  - size of the container
  - density of the gas
  - temperature of the gas
- The pressure of a sample of air in a manometer is 101.3 kPa. What is the partial pressure of nitrogen ( $N_2$ ) in the sample if the combined partial pressures of the other gases is 22.4 kPa?
  - 42.4 kPa
  - 78.9 kPa
  - 101.3 kPa
  - 123.7 kPa

Use the figure to answer the following questions.

- What instrument is illustrated in the figure? barometer
- Who invented this instrument? Evangelista Torricelli
- What are the two opposing forces that control the height of the mercury in the column? gravity and atmospheric pressure
- What does it mean when the level of mercury drops in the column? The atmospheric pressure has increased.



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# Chemistry Matter And Change Study Guide

**Darrell D. Ebbing**



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Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application *Study Guide and Solutions Manual, Fundamentals of General, Organic, and Biological Chemistry, Third Edition* John McMurry, Susan McMurry, 1999 Provides worked out solutions to text problems along with chapter by chapter outlines and a variety of self tests at the end of each chapter *Study Guide for The Anatomy and Physiology Learning System* Edith Applegate, 2010-03-30 Designed to accompany The Anatomy and Physiology Learning System 4th Edition by Edith Applegate this study guide helps you learn and review basic A P concepts Each chapter emphasizes medical terminology with a set of key terms word parts clinical terms and abbreviations and then adds a variety of fun filled learning exercises review questions a quiz and a word puzzle The study guide corresponds to the textbook chapter for chapter Chapter learning objectives help you focus on the most important material Key concepts are defined on the first page of each chapter in the workbook Learning exercises for each chapter include short answer matching and diagrams to label and color Self quizzes allow you to measure your progress and understanding Fun and Games features end each chapter with a variety of engaging puzzles covering words and concepts A chapter summary provides a brief review of each chapter A chapter review provides questions for reinforcement and review of the concepts in each chapter **Study Guide [to Accompany] General Chemistry** Dixie Goss, 2002 **Student Study Guide to accompany Chemistry: The Molecular Nature of Matter and Change** Martin Silberberg, 2008-01-14 This valuable ancillary is designed to help you recognize your learning style understand how to read classify a create a problem solving list and practice problem solving skills For each section of a chapter the author provides study objectives and a summary of the corresponding text Following the summary are sample problems with detailed solutions Each chapter has true false questions and a self test with all answers provided at the end of the chapter Chemistry, Study Guide James E. Brady, John R. Holum, 1993-02-11 Offers accurate lucid and interesting explanations of basic concepts and facts of chemistry while helping students develop skills in analytical thinking and problem solving Students are taught in a variety of ways to think of skills as tools that can be used to solve complex problems Several aids are included to help focus and inspire student interest frequent reference to common chemicals in commercial products numerous photographs of reactions in chapter practice exercises following worked examples

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