

# CHEMISTRY MULTIPLE CHOICE QUESTIONS

## A. Atoms, molecules and stoichiometry

2002 -2019



C

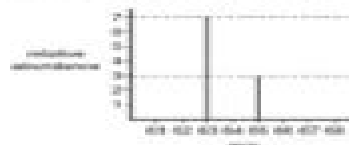
As a combustion, an alkane burns can be considered to have a fully set of 1.00 kg of methane (empirical formula  $\text{CH}_4$ ).

Which mass of carbon dioxide does a person produce each day if all the carbon they take in is digested and oxidised?

- A 0.007 kg    B 0.006 kg    C 1.30 kg    D 0.04 kg

(2002 NA/12)

The diagram shows the mass spectrum of a sample of naturally-occurring copper.



What is the relative atomic mass of this copper?

- A 63.2    B 63.5    C 63.8    D 64.0

(2002 NA/12)

Which pairs of compounds have the same empirical formula?

- ethane and ethene
- ethane and cyclohexane
- cyclohexane and oct-1-ene

(2002 NA/12)

A stream of  $10\text{cm}^3$  of methane and  $10\text{cm}^3$  of ethane was sparked with an excess of oxygen. After cooling to room temperature, the residual gas was passed through aqueous potassium hydroxide.

Which compounds have the empirical formula  $\text{CH}_2\text{O}$ ?

- methanol
- ethanoic acid
- methyl methanoate

(2002 ON/11)

What is the number of molecules in  $500\text{ cm}^3$  of oxygen under room conditions?

- A  $1.25 \times 10^{23}$   
 B  $1.25 \times 10^{24}$   
 C  $3.0 \times 10^{23}$   
 D  $3.0 \times 10^{24}$

(2004 MO/12)

In the preparation of soft margarine, glycerol trioleate



is fully hydrogenated so that, on average, one of its side-chains is converted into the  $\text{CH}_3(\text{CH}_2)_7\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$  moiety and two side-chains are converted into the  $\text{CH}_3(\text{CH}_2)_7\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$  moiety.

How many moles of hydrogen are required to convert one mole of glycerol trioleate into the soft margarine?

- A 4    B 5    C 6    D 8

(2008 MO/12)

Use of the Data Booklet is relevant to this question.

Analysed classes may obtain more credit awards of extra marks. Awarded to 10/17/2019

# Chemistry Multiple Choice Questions And Answers Doc

**SJ Ball**



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quantum theory relative abundance Rutherford model of atom shapes of orbitals solid iodine structure solids properties solvent extraction spectrometer spin quantum number states of matter stoichiometry sublimation thermometry scales types of solids unit cell van der Waals equation vapor pressure and spectrum

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