

MCQ Module – Understanding Chemical Equilibrium Practice

Instructions:

Answer the three multiple-choice questions below, choosing the correct response from the four options given. You will receive immediate feedback on your answers.

Question 1/3/3

1. What does the symbol K_c represent?

- a) Equilibrium constant for heterogeneous reactions
- b) Equilibrium constant for homogeneous reactions
- c) Equilibrium constant for heterogeneous reactions
- d) Equilibrium constant for homogeneous reactions

2. What does the symbol K_p represent in a reaction where all species are gaseous and the reaction is:

- a) $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{H}_2\text{O}(\text{g})$
- b) $\text{H}_2(\text{g}) + \text{Cl}_2(\text{g}) \rightleftharpoons 2\text{HCl}(\text{g})$
- c) $\text{H}_2(\text{g}) + \text{Cl}_2(\text{g}) \rightleftharpoons 2\text{HCl}(\text{l})$
- d) $\text{H}_2(\text{g}) + \text{Cl}_2(\text{g}) \rightleftharpoons 2\text{HCl}(\text{s})$

3. What are the units of the equilibrium constant for the reaction:

- a) mol^{-1}
- b) mol^{-2}
- c) $\text{mol}^{-1}\text{dm}^{-3}$
- d) $\text{mol}^{-2}\text{dm}^{-6}$

4. What does the symbol K_c represent in the reaction of water and steam:

- a) $\text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_2\text{O}(\text{g})$
- b) $\text{H}_2\text{O}(\text{g}) \rightleftharpoons \text{H}_2\text{O}(\text{l})$
- c) $\text{H}_2\text{O}(\text{g}) \rightleftharpoons \text{H}_2(\text{g}) + \text{O}_2(\text{g})$
- d) $\text{H}_2\text{O}(\text{g}) \rightleftharpoons \text{H}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g})$

5. What does the symbol K_p represent in the reaction of water and steam:

- a) $\text{H}_2\text{O}(\text{l})$
- b) $\text{H}_2\text{O}(\text{g})$
- c) $\text{H}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g})$
- d) $\text{H}_2(\text{g}) + \text{O}_2(\text{g})$

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