

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

## CHAPTER 9

## STUDY GUIDE FOR CONTENT MASTERY

## Covalent Bonding

### Section 11 The Covalent Bond

In your textbook, read about the nature of covalent bonds.

Use each of the terms below just once to complete the passage.

covalent bond	molecule	sigma bond	exothermic	$\pi$ bond
---------------	----------	------------	------------	------------

When sharing electrons occurs, the attachment between atoms that results is called a(n) **covalent bond**. When such an attachment is formed, bond dissociation energy is released, and the process is **exothermic**. When two or more atoms bond by means of electron sharing, the resulting particle is called a(n) **molecule**. If the electrons shared are contained between the two atoms, the attachment is called a(n) **sigma bond**. If the sharing involves the overlap of parallel orbitals, the attachment is called a(n)  **$\pi$  bond**.

In your textbook, read about single and multiple bonds and bond strengths.

Circle the letter of the choice that best completes the statement or answers the question.

- Is water from its elements such as hydrogen, oxygen, and oxygen normally **oxy**?
  - a single atom
  - a molecule containing three atoms
  - a molecule containing two atoms
  - a molecule containing four atoms
- How many electrons are shared in a double covalent bond?
  - none
  - two
  - four
  - eight
- Bond length is the distance between
  - two molecules of the same substance
  - the centers of two attached atoms
  - the electrons in two attached atoms
  - the orbitals of two attached atoms
- Which of the following relationships relating to bond length is generally correct?
  - As the number of bonds, the stronger the bond.
  - As the number of bonds, the weaker the bond.
  - As the number of bonds, the fewer the electrons in it.
  - As the number of bonds, the lower the bond dissociation energy.

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Chemistry: Matter and Change • Chapter 9

49

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## CHAPTER 9 STUDY GUIDE FOR CONTENT MASTERY

### Section 12 Naming Molecules

In your textbook, read about how binary compounds and acids are named from their formulas.

For each statement below, write *true* or *false*.

- false** 1. Binary molecular compounds are generally composed of a metal and a nonmetal.
- false** 2. The second element in the formula of a binary compound is named using the suffix *-ite*.
- false** 3. The prefix *mono-* indicates three atoms.
- true** 4. The prefix *hepta-* indicates six atoms.
- true** 5. In naming the first element in a formula, the prefix *mono-* is not used.
- true** 6. For binary acids, the hydrogen part of the compound is named using the prefix *hydro-*.
- false** 7. An oxyacid contains only two elements.
- false** 8. If the name of the cation of an oxyacid ends in *-ite*, the acid name contains the suffix *-ate*.

In your textbook, read about naming molecular compounds and oxyacids.

For each item in Column A, write the letter of the matching item in Column B.

## Column A

- c.**  $\text{N}_2\text{O}$       **b.**  $\text{Cl}_2$
- f.**  $\text{NO}$       **d.**  $\text{CO}_2$
- g.**  $\text{H}_2\text{SO}_4$       **e.**  $\text{H}_2\text{CO}_3$
- h.**  $\text{SO}_2$       **a.**  $\text{HNO}_3$
- b.**  $\text{H}_2\text{SO}_4$       **f.**  $\text{HNO}_2$
- d.**  $\text{HNO}_3$       **g.**  $\text{HNO}_2$
- a.**  $\text{HNO}_2$       **h.**  $\text{HNO}_3$

## Column B

- a.** hydrochloric acid
- b.** nitrogen trioxide
- c.** carbon monoxide
- d.** nitrous acid
- e.** carbonic acid
- f.** nitric acid
- g.** carbonic acid
- h.** nitric acid
- i.** carbon dioxide

50 Chemistry: Matter and Change • Chapter 9

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## **Table of Contents Chapter 7 Study Guide For Content Mastery Answer Key**

1. Understanding the eBook Chapter 7 Study Guide For Content Mastery Answer Key
  - The Rise of Digital Reading Chapter 7 Study Guide For Content Mastery Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 7 Study Guide For Content Mastery Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 7 Study Guide For Content Mastery Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 7 Study Guide For Content Mastery Answer Key
  - Personalized Recommendations

- Chapter 7 Study Guide For Content Mastery Answer Key User Reviews and Ratings
- Chapter 7 Study Guide For Content Mastery Answer Key and Bestseller Lists
- 5. Accessing Chapter 7 Study Guide For Content Mastery Answer Key Free and Paid eBooks
  - Chapter 7 Study Guide For Content Mastery Answer Key Public Domain eBooks
  - Chapter 7 Study Guide For Content Mastery Answer Key eBook Subscription Services
  - Chapter 7 Study Guide For Content Mastery Answer Key Budget-Friendly Options
- 6. Navigating Chapter 7 Study Guide For Content Mastery Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 7 Study Guide For Content Mastery Answer Key Compatibility with Devices
  - Chapter 7 Study Guide For Content Mastery Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 7 Study Guide For Content Mastery Answer Key
  - Highlighting and Note-Taking Chapter 7 Study Guide For Content Mastery Answer Key
  - Interactive Elements Chapter 7 Study Guide For Content Mastery Answer Key
- 8. Staying Engaged with Chapter 7 Study Guide For Content Mastery Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 7 Study Guide For Content Mastery Answer Key
- 9. Balancing eBooks and Physical Books Chapter 7 Study Guide For Content Mastery Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 7 Study Guide For Content Mastery Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 7 Study Guide For Content Mastery Answer Key
  - Setting Reading Goals Chapter 7 Study Guide For Content Mastery Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 7 Study Guide For Content Mastery Answer Key
  - Fact-Checking eBook Content of Chapter 7 Study Guide For Content Mastery Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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