

**KEY**

3. The pH in a 0.25 M solution of the acid HBrO is 4.65. Using this, calculate the value of  $K_a$  for the acid HBrO. (4 marks)  $\text{pH} = 4.65$  so  $[\text{H}_3\text{O}^+] = \text{antilog}(-4.65) = 2.239 \times 10^{-5} \text{ M}$

$$\text{HBrO} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{BrO}^-$$

[I]	0.25		0	0
[C]	$-2.239 \times 10^{-5}$		$+2.239 \times 10^{-5}$	$+2.239 \times 10^{-5}$
[E]	$0.25 - 2.239 \times 10^{-5}$		$2.239 \times 10^{-5}$	$2.239 \times 10^{-5}$

④

$$K_a = \frac{[\text{H}_3\text{O}^+][\text{BrO}^-]}{[\text{HBrO}]}$$

$$= \frac{(2.239 \times 10^{-5})^2}{(0.25 - 2.239 \times 10^{-5})} = 2.0 \times 10^{-9}$$

Answer  $K_a = 2.0 \times 10^{-9}$

4. The pH in a solution of benzoic acid is 2.355. Determine the molar concentration of the benzoic acid. (4 marks)  $\text{pH} = 2.355$   $[\text{H}_3\text{O}^+] = \text{antilog}(-2.355) = 0.004416 \text{ M}$

$$\text{C}_6\text{H}_5\text{COOH} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{C}_6\text{H}_5\text{COO}^-$$

[I]	$C_0$		0	0
[C]	$-0.004416$		$+0.004416$	$+0.004416$
[E]	$C_0 - 0.004416$		$0.004416$	$0.004416$

$$K_a = \frac{[\text{H}_3\text{O}^+][\text{C}_6\text{H}_5\text{COO}^-]}{[\text{C}_6\text{H}_5\text{COOH}]}$$

$$K_a = \frac{(0.004416)^2}{(C_0 - 0.004416)}$$

$$C_0 - 0.004416 = \frac{(0.004416)^2}{6.5 \times 10^{-5}} = 0.299976$$

$$C_0 = 0.299976 + 0.004416 = 0.30439 \text{ M}$$

Answer  $C_0 = 0.30 \text{ M}$

5. Find the value of  $K_b$  for the oxalate ion ( $\text{C}_2\text{O}_4^{2-}$ ). (1 mark)

$$K_b(\text{C}_2\text{O}_4^{2-}) = \frac{K_w}{K_a(\text{HC}_2\text{O}_4^-)} = \frac{1.00 \times 10^{-14}}{6.4 \times 10^{-5}} = 1.56 \times 10^{-10}$$

Answer  $K_b = 1.6 \times 10^{-10}$

6. The value of  $K_b$  for the weak base methylamine ( $\text{CH}_3\text{NH}_2$ ) is  $4.4 \times 10^{-4}$ . Calculate the value of  $K_a$  for the acid  $\text{CH}_3\text{NH}_3^+$ . (1 mark)

$$K_a(\text{CH}_3\text{NH}_3^+) = \frac{K_w}{K_b(\text{CH}_3\text{NH}_2)} = \frac{1.00 \times 10^{-14}}{4.4 \times 10^{-4}} = 2.27 \times 10^{-11}$$

Answer  $K_a = 2.3 \times 10^{-11}$

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# Basic Stoichiometry Phet Lab Answer Key

**Jicheng Xie**



**Basic Stoichiometry Phet Lab Answer Key :**

## **Basic Stoichiometry Phet Lab Answer Key** Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Basic Stoichiometry Phet Lab Answer Key**," published by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve to the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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