

Name: _____

Molecule Polarity PhET Lab

A study of electronegativity, bond polarity, and molecular polarity

Introduction:

In this atomic-level simulation, you will investigate how atoms' **electronegativity** value affects the bonds they produce. When two atoms bond, a pair of electrons is shared between atoms. Electronegativity is a measure of a single atom's ability to hoard electrons shared in that bond. In this lab you will work diligently, at your own pace, to answer a number of questions.

To begin, from what you've already learned about the protons and electrons in an atom, what would cause an atom to have a **high** electronegativity value?

Why might an atom have a **low** electronegativity value?

Procedure: **Two Atoms**

- ❖ Turn on (check) all view options. **GRAB** and spin the molecule with your mouse.
- ❖ Take your time and investigate how the binary compound's bond behaves when the atom's electronegativity and orientation are changed. Do not rush through this step.

Describe the bond formed between two atoms with **similar, low** electronegativities in terms of bond character.

Describe the bond formed between two atoms with **similar, high** electronegativities in terms of bond character.

Describe the bond formed between two atoms with **very different** electronegativities in terms of bond character.

Describe (in your own words) what is meant by **partial charges**, δ^- and δ^+ .

δ^- represents: _____

δ^+ represents: _____

What is **electron density**? _____

How does the density around a partial positive compare to the density of a partial negative?



Molecule Polarity

View

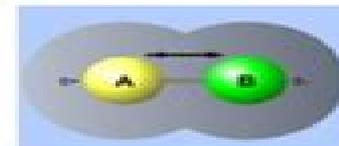
- Bond Dipole \longleftrightarrow
- Partial Charges
- Bond Character

Surface

- none
- Electrostatic Potential
- Electron Density

Electric Field

- on off



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