

CHEMISTRY 9TH - Solved Numerical

→ FUNDAMENTALS OF CHEMISTRY

CHAPTER# 01

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1. Calculate the number of moles in 30 gm of H_3PO_4 .

Solution:

Data: Mass of H_3PO_4 = 30 gm

$$\begin{aligned} \text{Molecular mass of } H_3PO_4 &= 1(3) + 30.97 + 16(4) \\ &= 3 + 30.97 + 64 \\ &= 97.97 \text{ or } 98 \end{aligned}$$

Calculation:

$$\text{Number of Moles} = \frac{30}{98} = \boxed{0.3061 \text{ moles}} \text{ Ans}$$

2. Calculate the number of moles, in 50 g of each:

(i) Na (ii) H_2O

$$\text{Number of Moles} = \frac{\text{Given mass of substance}}{\text{Atomic mass or Formula mass}}$$

(i) Given,

(i) Number of moles = ?

(ii) Given mass of Na = 50 g

(iii) Atomic mass of Na = 23 a.m.u.

$$\Rightarrow \text{No of moles of Na} = \frac{50}{23} = 2.173 \text{ moles (of Na)}$$

(ii) Given,

(i) Number of moles of H_2O = ?

(ii) Formula mass of H_2O = 18 a.m.u.

(iii) Given mass of H_2O = 50 g

$$= \text{No of moles of } H_2O = \frac{50}{18} = 2.777 \text{ moles (of } H_2O)$$

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Chao Zhang



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Table of Contents 9th Class Chemistry Notes Unit 1 Fundamentals Of

1. Understanding the eBook 9th Class Chemistry Notes Unit 1 Fundamentals Of
 - The Rise of Digital Reading 9th Class Chemistry Notes Unit 1 Fundamentals Of
 - Advantages of eBooks Over Traditional Books
2. Identifying 9th Class Chemistry Notes Unit 1 Fundamentals Of
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 9th Class Chemistry Notes Unit 1 Fundamentals Of
 - User-Friendly Interface
4. Exploring eBook Recommendations from 9th Class Chemistry Notes Unit 1 Fundamentals Of
 - Personalized Recommendations
 - 9th Class Chemistry Notes Unit 1 Fundamentals Of User Reviews and Ratings

- 9th Class Chemistry Notes Unit 1 Fundamentals Of and Bestseller Lists
- 5. Accessing 9th Class Chemistry Notes Unit 1 Fundamentals Of Free and Paid eBooks
 - 9th Class Chemistry Notes Unit 1 Fundamentals Of Public Domain eBooks
 - 9th Class Chemistry Notes Unit 1 Fundamentals Of eBook Subscription Services
 - 9th Class Chemistry Notes Unit 1 Fundamentals Of Budget-Friendly Options
- 6. Navigating 9th Class Chemistry Notes Unit 1 Fundamentals Of eBook Formats
 - ePub, PDF, MOBI, and More
 - 9th Class Chemistry Notes Unit 1 Fundamentals Of Compatibility with Devices
 - 9th Class Chemistry Notes Unit 1 Fundamentals Of Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 9th Class Chemistry Notes Unit 1 Fundamentals Of
 - Highlighting and Note-Taking 9th Class Chemistry Notes Unit 1 Fundamentals Of
 - Interactive Elements 9th Class Chemistry Notes Unit 1 Fundamentals Of
- 8. Staying Engaged with 9th Class Chemistry Notes Unit 1 Fundamentals Of
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 9th Class Chemistry Notes Unit 1 Fundamentals Of
- 9. Balancing eBooks and Physical Books 9th Class Chemistry Notes Unit 1 Fundamentals Of
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 9th Class Chemistry Notes Unit 1 Fundamentals Of
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 9th Class Chemistry Notes Unit 1 Fundamentals Of
 - Setting Reading Goals 9th Class Chemistry Notes Unit 1 Fundamentals Of
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 9th Class Chemistry Notes Unit 1 Fundamentals Of
 - Fact-Checking eBook Content of 9th Class Chemistry Notes Unit 1 Fundamentals Of
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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