

# K<sub>w</sub> CALCULATION

$$K_w = [H^+][OH^-]$$



$$\Rightarrow [H^+] = \frac{K_w}{[OH^-]} = \frac{1 \times 10^{-14}}{0.46}$$

$$pH = -\log[H^+]$$

$$= -\log(2.17391 \times 10^{-14})$$

$$= 13.66$$

pH to 2 d.p.!

A-LEVEL CHEMISTRY

# Calculations For A Level Chemistry

**Colorado State University**



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