



COVALENT BONDING

Practice Problems

Do your answers to the following problems.

SECTION 8.1 MOLECULAR COMPOUNDS

- Classify each of the following as either ionic or covalent.
 - NaCl
 - H_2O
 - HCl
 - KNO_3
 - H_2SO_4
 - H_2O_2
- Write the formula for the following compounds:
 - hydrogen chloride
 - hydrogen sulfide
 - hydrogen cyanide
 - hydrogen peroxide
 - hydrogen selenide
 - hydrogen telluride
- Write the formula for the following compounds:
 - hydrogen bromide
 - hydrogen iodide
 - hydrogen selenide
 - hydrogen telluride
 - hydrogen phosphide
 - hydrogen arsenide
- Write the formula for the following compounds:
 - hydrogen fluoride
 - hydrogen chloride
 - hydrogen bromide
 - hydrogen iodide
 - hydrogen selenide
 - hydrogen telluride

SECTION 8.2 THE NATURE OF COVALENT BONDING

- Draw the Lewis structure for hydrogen fluoride. $\text{H}-\text{F}$
- Draw the Lewis structure for hydrogen chloride. $\text{H}-\text{Cl}$
- Draw the Lewis structure for hydrogen bromide. $\text{H}-\text{Br}$
- Draw the Lewis structure for hydrogen iodide. $\text{H}-\text{I}$
- Draw the Lewis structure for hydrogen cyanide. $\text{H}-\text{C}\equiv\text{N}$
- Draw the Lewis structure for hydrogen sulfide. $\text{H}-\text{S}-\text{H}$
- Draw the Lewis structure for hydrogen selenide. $\text{H}-\text{Se}-\text{H}$
- Draw the Lewis structure for hydrogen telluride. $\text{H}-\text{Te}-\text{H}$
- Draw the Lewis structure for hydrogen phosphide. $\text{H}-\text{P}-\text{H}$
- Draw the Lewis structure for hydrogen arsenide. $\text{H}-\text{As}-\text{H}$



SECTION 8.3 BONDING IN THE LHS

- Write the Lewis structure for hydrogen cyanide. $\text{H}-\text{C}\equiv\text{N}$
- Write the Lewis structure for hydrogen sulfide. $\text{H}-\text{S}-\text{H}$

Chapter 8 Covalent Bonding Practice Problems Answers Bing

Mike Jess



Chapter 8 Covalent Bonding Practice Problems Answers Bing:

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