



COVALENT BONDING

Practice Problems

Do your answers to the following problems.

SECTION 8.1 MOLECULAR COMPOUNDS

- Classify each of the following as either ionic or molecular.
 - NaCl
 - H_2O
 - HCl
 - KNO_3
 - H_2SO_4
 - H_2O_2
- Write the formula for the following molecular compounds:
 - nitrogen dioxide
 - carbon dioxide
 - carbon monoxide
 - nitrogen monoxide
 - nitrogen trioxide
 - nitrogen pentoxide

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 - nitrogen monoxide
 - nitrogen trioxide
 - nitrogen pentoxide

SECTION 8.2 THE NATURE OF COVALENT BONDING

- Show covalent bonding in hydrogen fluoride: $\text{H}-\text{F}$
- Show covalent bonding in methane: CH_4
- Show covalent bonding in carbon dioxide: $\text{O}=\text{C}=\text{O}$
- Show covalent bonding in acetylene: $\text{H}-\text{C}\equiv\text{C}-\text{H}$
- Draw Lewis structures for the following:
 - H_2O
 - CO_2
 - NH_3
 - CH_4
 - H_2O_2
 - C_2H_2
 - C_2H_4
 - C_2H_6



SECTION 8.3 BONDING: THE ORBITALS

- Which orbital and sub-orbital do you combine to form the sigma bond in H_2 ? The sigma bond in H_2 is formed from two 1s orbitals.
- Which orbital and sub-orbital do you combine to form the sigma bond in H_2 ? The sigma bond in H_2 is formed from two 1s orbitals.

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Daniel F McAuley



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