

CHEMISTRY REVIEW: KEY CHAPTER 9 – CHEMICAL NAMES AND FORMULAS

- In a Stock name such as iron III sulfate, the roman numeral tells us **C**
 - How many atoms of Fe are in one formula unit
 - How many sulfate ions can be attached to the iron atom
 - The charge on each Fe ion
 - The total positive charge of the formula unit
- The result of changing a subscript in a correctly written chemical formula is to **c**
 - Change the number of moles represented in the formula
 - Change the charges on the other ions in the compound
 - Change the formula so that it no longer represents the compound
 - Have no effect on the formula
- Derive a generalization for when an acid name will end in the suffix **-ic** or **-ous**
If ending is -ic, the anion ended in -ate; if ending is -ous, the anion ended in -ite
- Derive a generalization for when an acid name will begin with the prefix **hydro-**.
If prefix is hydro-, the anion ended in -ide. This is most likely a monatomic ion.
- Fill in the blanks in the table below

| Compound Name | Formula |
|-----------------------------|----------------|
| Aluminum sulfide | Al_2S_3 |
| Aluminum sulfate | $Al_2(SO_4)_3$ |
| Lead II chloride | $PbCl_2$ |
| Ammonium phosphite | $(NH_4)_2PO_3$ |
| Hydroiodic acid | HCl |
| Dinitrogen pentoxide | N_2O_5 |
| Copper II carbonate | $CuCO_3$ |
| Sodium sulfide | Na_2S |
| Ammonium phosphate | $(NH_4)_3PO_4$ |
| Tin IV sulfite | $Sn(SO_3)_2$ |
| Nitrous acid | HNO_2 |
| Magnesium perchlorate | $Mg(ClO_4)_2$ |
| Iron II nitrate | $Fe(NO_3)_2$ |
| Iron III nitrate | $Fe(NO_3)_3$ |
| Cobalt II oxide | CoO |
| Sulfur trioxide | SO_3 |
| Boron pentafluoride | BF_5 |
| Tetraphosphorus octasulfide | P_4S_8 |
| Tetraiodide nonoxide | I_4O_9 |

- Name the following compound and identify as molecular, ionic or acid
 - $Ca(HCO_3)_2$ **calcium bicarbonate - I**
 - $Fe(ClO_2)_3$ **Iron III chlorite - I**
 - $Fe(ClO_4)_2$ **Iron II chlorate - I**
 - N_2O_5 **dinitrogen pentoxide - M**

Chapter 9 Chemical Names Formulas Answers Key

RJ Shavelson



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