

## Prentice Hall Physical Science Chapter 4 Atomic Structure

### 4.1 Studying Atoms

- A. **Ancient Greek Models of Atoms**
  - Democritus believed all matter was made of extremely small particles that could not be divided
  - he called the particles atoms (from the Greek atomos meaning indivisible)
  - he thought there were different types of atoms with different properties (liquids had round, smooth atoms and solids were rough and prickly)
- B. **Dalton's Atomic Theory**
  - he proved atoms exist by discovering that compounds have a fixed composition
  - Dalton's Theory:
    - all elements are composed of atoms
    - all atoms of the same element have the same mass, and atoms of different elements have different masses
    - compounds contain atoms of more than one element
    - in a particular compound, atoms of different element always combine in the same way
  - he pictured atoms as solid spheres
  - in time it was found that not all of this was completely correct
- C. **Thomson's Model of the Atom**
  - some objects have a positive or negative charge and objects with the same charge repel and objects with opposite charges attract
  - Thomson used an electric current to learn more about atoms
  - he passed an electric current through a gas and a glowing beam appeared
  - the beam was repelled by a negatively charged metal plate (like charges repel and opposite charges attract)
  - Thomson concluded that the beam must be made of negatively charged particles, smaller than atoms.
  - since the atom was known to be electrically neutral, he developed a model in which negative charges were evenly scattered throughout an atom filled with a positively charged pudding like material
  - called the "Plum Pudding" model (or chocolate chip cookie dough)
- D. **Rutherford's Atomic Theory**
  - Rutherford discovered positively charged, fast moving particles called alpha particles
  - he shot some of these particles a piece of gold foil, thinking they would go straight through
  - when he actually performed the experiment, he found that the positively charged alpha particles were greatly deflected by the gold (figure 7, p. 104)
  - he determined that the alpha particles must have come very close to another positively charged particle causing them to be repelled and put off course
  - this led him to believe that the positive charge in an atom is not evenly distributed, but concentrated in a very small central area he called the nucleus
  - his model of the atom is a small, positively charged nucleus surrounded by negatively charged particles

# Chapter Assessment Prentice Hall Physical Science

**Michael J. Padilla, David V. Frank, John  
G. Little, Steve Miller**

## **Chapter Assessment Prentice Hall Physical Science:**

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*Resources for Teaching Middle School Science* Smithsonian Institution, National Academy of Engineering, National Science Resources Center of the National Academy of Sciences, Institute of Medicine, 1998-04-30 With age appropriate inquiry centered curriculum materials and sound teaching practices middle school science can capture the interest and energy of adolescent students and expand their understanding of the world around them *Resources for Teaching Middle School Science* developed by the National Science Resources Center NSRC is a valuable tool for identifying and selecting effective science curriculum materials that will engage students in grades 6 through 8 The volume describes more than 400 curriculum titles that are aligned with the National Science Education Standards This completely new guide follows on the success of *Resources for Teaching Elementary School Science* the first in the NSRC series of annotated guides to hands on inquiry centered curriculum materials and other resources for science teachers The curriculum materials in the new guide are grouped in five chapters by scientific area Physical Science Life Science Environmental Science Earth and Space Science and Multidisciplinary and Applied Science They are also grouped by type core materials supplementary units and science activity books Each annotation of curriculum material includes a recommended grade level a description of the activities involved and of what students can be expected to learn a list of accompanying materials a reading level and ordering information The curriculum materials included in this book were selected by panels of teachers and scientists using evaluation criteria developed for the guide The criteria reflect and incorporate goals and principles of the National Science Education Standards The annotations designate the specific content standards on which these curriculum pieces focus In addition to the curriculum chapters the guide contains six chapters of diverse resources that are directly relevant to middle school science Among these is a chapter on educational software and multimedia programs chapters on books about science and teaching directories and guides to science trade books and periodicals for teachers and students Another section features institutional resources One chapter lists about 600 science centers museums and zoos where teachers can take middle school students for interactive science experiences Another chapter describes nearly 140 professional associations and U S government agencies that offer resources and assistance Authoritative extensive and thoroughly indexed and the only guide of its kind *Resources for Teaching Middle School Science* will be the most used book on the shelf for science teachers school administrators teacher trainers science curriculum specialists advocates of hands on science teaching and

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## **Table of Contents Chapter Assessment Prentice Hall Physical Science**

1. Understanding the eBook Chapter Assessment Prentice Hall Physical Science
  - The Rise of Digital Reading Chapter Assessment Prentice Hall Physical Science
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter Assessment Prentice Hall Physical Science
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter Assessment Prentice Hall Physical Science
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter Assessment Prentice Hall Physical Science
  - Personalized Recommendations
  - Chapter Assessment Prentice Hall Physical Science User Reviews and Ratings
  - Chapter Assessment Prentice Hall Physical Science and Bestseller Lists
5. Accessing Chapter Assessment Prentice Hall Physical Science Free and Paid eBooks
  - Chapter Assessment Prentice Hall Physical Science Public Domain eBooks
  - Chapter Assessment Prentice Hall Physical Science eBook Subscription Services

- Chapter Assessment Prentice Hall Physical Science Budget-Friendly Options
- 6. Navigating Chapter Assessment Prentice Hall Physical Science eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter Assessment Prentice Hall Physical Science Compatibility with Devices
  - Chapter Assessment Prentice Hall Physical Science Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter Assessment Prentice Hall Physical Science
  - Highlighting and Note-Taking Chapter Assessment Prentice Hall Physical Science
  - Interactive Elements Chapter Assessment Prentice Hall Physical Science
- 8. Staying Engaged with Chapter Assessment Prentice Hall Physical Science
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter Assessment Prentice Hall Physical Science
- 9. Balancing eBooks and Physical Books Chapter Assessment Prentice Hall Physical Science
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter Assessment Prentice Hall Physical Science
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter Assessment Prentice Hall Physical Science
  - Setting Reading Goals Chapter Assessment Prentice Hall Physical Science
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter Assessment Prentice Hall Physical Science
  - Fact-Checking eBook Content of Chapter Assessment Prentice Hall Physical Science
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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