

# General Chemistry 11

## Activity Sheet 1 for Quarter 1 Week 3 (Paired/Individual Activity)

(TUESDAY)

**Reminders: Accomplish the activities every Tuesday and Thursday during your schedule of Chemistry class. Submit all answers in the google classwork.**

### Activity 1. What a Molar Mass!

**Activity 1.a. Calculate the molar mass of the following compounds:**

Compound	Solution	Molar Mass
$\text{CO}_2$	$\begin{array}{r} \text{C} = 12 \\ \text{O} = 16 \\ 12 \times 1 = 12 \\ 16 \times 2 = 32 \\ \hline 44\text{g} \end{array}$	44g/mol
$\text{Fe}_3(\text{PO}_4)_2$	$\begin{array}{r} \text{Fe} = 56 \\ \text{P} = 31 \\ \text{O} = 16 \\ 56 \times 3 = 168 \\ 31 \times 2 = 62 \\ 16 \times 4 = 64 \times 2 = 128 \\ \hline 358\text{g} \end{array}$	358g/mol
$\text{CH}_3\text{COOH}$	$\begin{array}{r} \text{C} = 12 \\ \text{H} = 1 \\ \text{O} = 16 \\ 12 \times 2 = 24 \\ 1 \times 3 + 1 = 4 \\ 16 \times 2 = 32 \\ \hline 60\text{g} \end{array}$	60g/mol
$\text{H}_2\text{O}$	$\begin{array}{r} \text{H} = 1 \\ \text{O} = 16 \\ 1 \times 2 = 2 \\ 16 \times 1 = 16 \\ \hline 18\text{g} \end{array}$	18g/mol
$\text{CaSO}_4$	$\begin{array}{r} \text{Ca} = 40 \\ \text{S} = 32 \\ \text{O} = 16 \\ 40 \times 1 = 40 \\ 32 \times 1 = 32 \\ 16 \times 4 = 64 \\ \hline 136\text{g} \end{array}$	136g/mol

**Activity 1.b. Direction: Identify whether the following is an empirical or molecular formula:**

a.  $\text{CH}$  Empirical Formula

d.  $\text{CH}_2\text{O}$  Molecular Formula

b.  $\text{CO}_2$  Empirical Formula

e.  $\text{C}_3\text{H}_4\text{O}_3$  Empirical Formula

c.  $\text{AlCl}_3$  Molecular Formula

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**Ying Liu**



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