

**WORKSHEET**  
**CHEMISTRY FOR ENGINEERING STUDENTS, LECTURE 1**  
 Molecular formula of hydrocarbon or a mixture of hydrocarbons

**ANSWERS KEY**

**QUESTION 1. Calculate the empirical formula of hydrocarbon and its molecular formula.**

1. A hydrocarbon of 0.100 g is completely burnt to give 0.345 g carbon dioxide and 0.070 g water. Calculate the empirical formula of the hydrocarbon.
  - (a) write the balanced chemical equation for the reaction  
 $C_xH_y + O_2 \rightarrow CO_2 + H_2O$
  - (b) calculate the relative atomic weights of carbon and hydrogen as they are given in the periodic table
  - (c) calculate the mass of the hydrocarbon burnt and of carbon, H and O  
 $0.100 \text{ g}$
  
2. A 0.100 g hydrocarbon is burnt to produce 0.345 g carbon dioxide and 0.070 g water. Calculate the empirical formula of the hydrocarbon.
  - (a) write the balanced chemical equation for the reaction  
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  - (b) calculate the relative atomic weights of carbon and hydrogen as they are given in the periodic table
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3. A hydrocarbon of 0.100 g is burnt to give 0.345 g carbon dioxide and 0.070 g water. Calculate the empirical formula of the hydrocarbon.
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 $C_xH_y + O_2 \rightarrow CO_2 + H_2O$
  - (b) calculate the mass of the hydrocarbon burnt and of carbon, H and O  
 $0.100 \text{ g}$
  - (c) calculate the relative atomic weights of carbon and hydrogen as they are given in the periodic table

**QUESTION 2. Calculate the molecular formula of the hydrocarbon.**

1. The molar mass of a hydrocarbon is 58 g mol<sup>-1</sup>. Calculate the molecular formula of the hydrocarbon.
  - (a) write the balanced chemical equation for the reaction  
 $C_xH_y + O_2 \rightarrow CO_2 + H_2O$
  - (b) calculate the mass of the hydrocarbon burnt and of carbon, H and O  
 $0.100 \text{ g}$

# Chem 21 Lab Answers

**Henry Dorin**



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