

CHEMISTRY 11 – WORD EQUATIONS AND TYPES OF CHEMICAL REACTIONS WORKSHEET

1) Translate the following word equations into chemical symbols and balance the resulting equations. Do not include the phases.

(a) Potassium + water → potassium hydroxide + hydrogen



(b) Nitrogen + hydrogen → ammonia



2) In each of the following write a balanced chemical equation, including the phases.

(a) Liquid phosphoric acid reacts with aqueous barium hydroxide to give water and a precipitate of barium phosphate



(b) Solid aluminium oxide and aqueous sulphuric acid produce water and aqueous aluminium sulphate



(c) Nitrogen trifluoride gas and hydrogen gas react to form nitrogen gas and gaseous hydrogen fluoride



(d) Powdered sodium carbonate and aqueous hydrobromic acid react to form carbon dioxide gas, aqueous sodium bromide and water



3) Classify the following reactions.

Reaction	Type
$2 \text{C}_2\text{H}_2 + 5 \text{O}_2 \rightarrow 4 \text{CO}_2 + 2 \text{H}_2\text{O}$	Combustion
$\text{Mg} + \text{CuSO}_4 \rightarrow \text{MgSO}_4 + \text{Cu}$	Single Replacement
$4 \text{Na} + \text{O}_2 \rightarrow 2 \text{Na}_2\text{O}$	Synthesis
$\text{Sn}(\text{OH})_4 + 4 \text{HBr} \rightarrow 4 \text{H}_2\text{O} + \text{SnBr}_4$	Neutralization
$2 \text{Fe}(\text{NO}_3)_3 + 3 \text{MgS} \rightarrow \text{Fe}_2\text{S}_3 + 3 \text{Mg}(\text{NO}_3)_2$	Double Replacement
$2 \text{N}_2\text{O} \rightarrow 2 \text{N}_2 + \text{O}_2$	Decomposition

Chemical Equations And Reactions 8 2 Answers

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Table of Contents Chemical Equations And Reactions 8 2 Answers

1. Understanding the eBook Chemical Equations And Reactions 8 2 Answers
 - The Rise of Digital Reading Chemical Equations And Reactions 8 2 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Equations And Reactions 8 2 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemical Equations And Reactions 8 2 Answers
 - User-Friendly Interface

4. Exploring eBook Recommendations from Chemical Equations And Reactions 8 2 Answers
 - Personalized Recommendations
 - Chemical Equations And Reactions 8 2 Answers User Reviews and Ratings
 - Chemical Equations And Reactions 8 2 Answers and Bestseller Lists
5. Accessing Chemical Equations And Reactions 8 2 Answers Free and Paid eBooks
 - Chemical Equations And Reactions 8 2 Answers Public Domain eBooks
 - Chemical Equations And Reactions 8 2 Answers eBook Subscription Services
 - Chemical Equations And Reactions 8 2 Answers Budget-Friendly Options
6. Navigating Chemical Equations And Reactions 8 2 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemical Equations And Reactions 8 2 Answers Compatibility with Devices
 - Chemical Equations And Reactions 8 2 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemical Equations And Reactions 8 2 Answers
 - Highlighting and Note-Taking Chemical Equations And Reactions 8 2 Answers
 - Interactive Elements Chemical Equations And Reactions 8 2 Answers
8. Staying Engaged with Chemical Equations And Reactions 8 2 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemical Equations And Reactions 8 2 Answers
9. Balancing eBooks and Physical Books Chemical Equations And Reactions 8 2 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemical Equations And Reactions 8 2 Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemical Equations And Reactions 8 2 Answers
 - Setting Reading Goals Chemical Equations And Reactions 8 2 Answers
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Chemical Equations And Reactions 8 2 Answers
 - Fact-Checking eBook Content of Chemical Equations And Reactions 8 2 Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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