

# Stoichiometry & the mole notes

## the mole concept

the mole (mol) is the SI unit for measuring the amount of a substance.

avogadro's number:

$$1 \text{ mol} = 6.022 \times 10^{23} \text{ particles}$$

(particles can be atoms, molecules, or formula units)

molar mass:

the mass of one mole of a substance in grams.

found by adding the atomic masses from the periodic table

example

$$\text{H}_2\text{O} (2 \cdot 1.008 \text{ g/mol}) + (1 \cdot 16.00 \text{ g/mol}) = 18.016 \text{ g/mol}$$

## Example Problems

1. how many atoms are in 3.5 moles of Fe?

$$3.5 \text{ mol Fe} \cdot \frac{6.022 \cdot 10^{23} \text{ atoms}}{1 \text{ mol}} = 2.11 \cdot 10^{24} \text{ atoms}$$

2. what is the mass of 2.00 moles of NaCl?

find molar mass:

$$\text{Na} : 22.99 \text{ g/mol}$$

$$\text{Cl} : 35.45 \text{ g/mol}$$

$$\text{Total} : 58.44 \text{ g/mol}$$

convert

$$2.00 \text{ mol} \cdot \frac{58.44 \text{ g}}{1 \text{ mol}} = 116.88 \text{ g}$$



## Conversions Involving the mole

conversion

particles  $\leftrightarrow$  moles

moles  $\leftrightarrow$  mass

moles  $\leftrightarrow$  volume (L) (at STP)

factor to use

$$1 \text{ mol} = 6.022 \cdot 10^{23} \text{ particles}$$

$$1 \text{ mol} = \text{molar mass (g)}$$

$$1 \text{ mol} = 22.4 \text{ L of gas}$$

## Stoichiometry: the study of quantitative relationships in chemical reactions

Balanced chemical equations show the mole ratio of reactants & products

mole ratios come from coefficients in a balanced equation.

Stoichiometry steps:

1. Balance the equation

2. convert given quantity into moles (if necessary)

3. Use mole ratio from balanced equation.

4. Convert moles back to desired unit (grams, liters, etc.)

Example:



question: how many grams of  $\text{H}_2\text{O}$  are produced from 5.00 moles of

$\text{O}_2$ ? Step 1: Use mole ratio

$$5.00 \text{ mol O}_2 \cdot \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol O}_2} = 10.0 \text{ mol H}_2\text{O}$$

Step 2: Convert to grams

$$10.0 \text{ mol H}_2\text{O} \cdot \frac{18.016 \text{ g}}{1 \text{ mol}} = 180.16 \text{ g H}_2\text{O}$$

## Limiting Reactant & Percent Yield

limiting reactant (LR): the reactant that runs out first, limiting product formation

excess reactant: the reactant that remains after the reaction is complete

Steps to identify LR

1. Convert reactants to moles

2. Use mole ratios to do

# Chemistry Notes Chapter 9 Stoichiometry

**Clifford Lane**



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## Table of Contents Chemistry Notes Chapter 9 Stoichiometry

1. Understanding the eBook Chemistry Notes Chapter 9 Stoichiometry
  - The Rise of Digital Reading Chemistry Notes Chapter 9 Stoichiometry
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Notes Chapter 9 Stoichiometry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an eBook Chemistry Notes Chapter 9 Stoichiometry
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Notes Chapter 9 Stoichiometry
  - Personalized Recommendations
  - Chemistry Notes Chapter 9 Stoichiometry User Reviews and Ratings
  - Chemistry Notes Chapter 9 Stoichiometry and Bestseller Lists
5. Accessing Chemistry Notes Chapter 9 Stoichiometry Free and Paid eBooks
  - Chemistry Notes Chapter 9 Stoichiometry Public Domain eBooks
  - Chemistry Notes Chapter 9 Stoichiometry eBook Subscription Services
  - Chemistry Notes Chapter 9 Stoichiometry Budget-Friendly Options
6. Navigating Chemistry Notes Chapter 9 Stoichiometry eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Notes Chapter 9 Stoichiometry Compatibility with Devices
  - Chemistry Notes Chapter 9 Stoichiometry Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Notes Chapter 9 Stoichiometry
  - Highlighting and Note-Taking Chemistry Notes Chapter 9 Stoichiometry
  - Interactive Elements Chemistry Notes Chapter 9 Stoichiometry

8. Staying Engaged with Chemistry Notes Chapter 9 Stoichiometry
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Notes Chapter 9 Stoichiometry
9. Balancing eBooks and Physical Books Chemistry Notes Chapter 9 Stoichiometry
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Notes Chapter 9 Stoichiometry
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Notes Chapter 9 Stoichiometry
  - Setting Reading Goals Chemistry Notes Chapter 9 Stoichiometry
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Notes Chapter 9 Stoichiometry
  - Fact-Checking eBook Content of Chemistry Notes Chapter 9 Stoichiometry
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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